

Topic 4 Mixtures and Solutions

4.1 Miscibility and Solutions

Solvents can be considered as polar (e.g. water, methanol) or non-polar (e.g. hexane, turpentine, petrol).

- Identify water as a polar solvent and hydrocarbons as non-polar solvents.

Polar and nonpolar solvents do not readily mix.

- Identify a solvent as polar or non-polar based on its miscibility with water and hydrocarbons.

Highly polar molecular substances are more soluble in water than nonpolar molecules of a similar size.

Molecular substances with small molecules are more soluble in water than larger molecules of similar polarity.

- Predict, given the structural formulae, which of two compounds would be more soluble in polar and non-polar solvents.

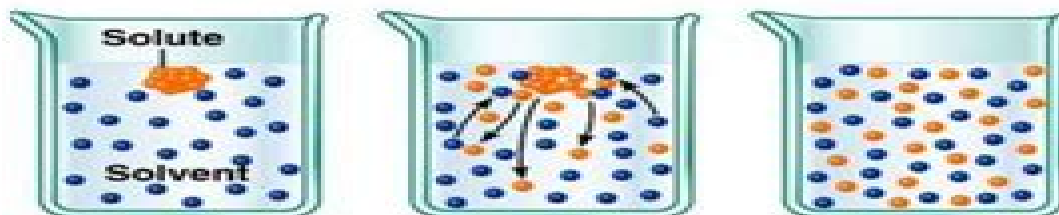
Compounds with non-polar and polar or ionic components facilitate the mixing of polar and non-polar substances.

A solution is a mixture in which a solute is dissolved in a solvent. The solute and solvent cannot be distinguished from one another and the bulk (e.g. colour, pH, conductivity etc.) properties of the solution are consistent throughout (when mixed).

During dissolution, the following processes occur:

- Solute particles are separated from one another
- Solvent particles are separated from one another
- Solute and solvent particles are attracted to one another

Note miscibility refers to the solubility of one liquid in another.



The ability of a solute to dissolve will be determined, in part, by the polarity of the its particles and the polarity of the solvent's particles. As a rule of thumb, we say "like dissolves like".

However, it is more correct to suggest that a solute of similar polarity to a solvent will dissolve in a solvent should size allow it.

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Patrick Vollmar



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