



PRE-LAB QUESTIONS

1. What would you estimate the pH of hydrochloric acid, a strong acid, to be? What is the pH of water?
2. Use the Henderson-Hasselbalch equation to estimate the pH of a buffer solution that is composed of 20 mL of 10 M sodium formate and 20 mL of 1 M formic acid (Note: the pK_a of formic acid is 3.75).
3. Look at Table 4 in the procedure portion of the experiment. Calculate the pH you would expect each of the buffer solutions (A, B, C, D, and E) to be using the Henderson-Hasselbalch equation, assuming that the solutions of acetic acid and sodium acetate are equimolar.



EXPERIMENT 1: PREPARING A BUFFER

Data Sheet

Table 3: Sodium Acetate Data

Sodium Acetate (g)	Molarity of Sodium Acetate (Mol/L)
4.0 g	0.4784

Table 4: Buffer Solutions and pH Readings for Beakers A, B, C, D, and E

Buffer	mL of Acetic Acid	mL of Sodium Acetate	pH measured
A	5	5	4.79
B	5	1	4.12
C	10	1	3.91
D	1	10	5.73
E	1	5	5.42

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