

- 4) $n = 3 ; l = 0 ; m = 0 ; s = +1/2$
 5) $n = 3 ; l = 1 ; m = -1 ; s = +1/2$

Exercice 4 :

Indiquez en le justifiant le nombre d'éléments présents dans les 2ème, 3ème, 4ème, 5ème et 6ème périodes de la classification.

Exercice 5 :

Un élément de numéro atomique inférieur à 20 possède un électron célibataire. Quelles sont les diverses possibilités ?

On sait de plus que cet électron célibataire est unique. Quelles sont les diverses possibilités ?

Cet élément appartient à la période de l'argon. Quelles sont les diverses possibilités ?

Cet élément appartient au groupe du Francium ($Z = 86$). Quel est cet élément ?

CORRIGE

ATOMISTIQUE

Exercice 1 :

Pb $Z = 82$ et $M = 207,2$

$$M = 207,2 \text{ g mol}^{-1} \Leftrightarrow A = 207 \Rightarrow N = A - Z = 207 - 82 = 125$$

Masse des électrons :

$$m_{\text{électrons}} = Z \cdot m_e = 82 \cdot 9,1095 \cdot 10^{-31} = 7,5 \cdot 10^{-29} \text{ kg} = 7,5 \cdot 10^{-28} \text{ g}$$

Masse de l'atome :

$$M = 207,2 \text{ g mol}^{-1} \Leftrightarrow m_{\text{atome}} = 207,2 \text{ u.m.a}$$

$$1 \text{ u.m.a} = 1 / N \text{ g} = 1 / 6,022 \cdot 10^{23} = 1,66 \cdot 10^{-24} \text{ g}$$

$$m_{\text{atome}} = 207,2 \cdot 1,66 \cdot 10^{-24} = 3,4 \cdot 10^{-22} \text{ g}$$

$$m_{\text{atome}} / m_{\text{électrons}} \approx 4586$$

H $Z = 1$ et $M = 1,008$

$$M = 1,008 \text{ g mol}^{-1} \Leftrightarrow A = 1 \Rightarrow N = A - Z = 1 - 1 = 0$$

Masse des électrons :

$$m_{\text{électrons}} = Z \cdot m_e = 1 \cdot 9,1095 \cdot 10^{-31} \approx 9,1 \cdot 10^{-31} \text{ kg} = 9,1 \cdot 10^{-28} \text{ g}$$

Masse de l'atome :

$$M = 1,008 \text{ g mol}^{-1} \Leftrightarrow m_{\text{atome}} = 1,008 \text{ u.m.a}$$

$$1 \text{ u.m.a} = 1 / N \text{ g} = 1 / 6,022 \cdot 10^{23} = 1,66 \cdot 10^{-24} \text{ g}$$

$$m_{\text{atome}} = 1,008 \cdot 1,66 \cdot 10^{-24} \approx 1,67 \cdot 10^{-24} \text{ g}$$

$$m_{\text{atome}} / m_{\text{électrons}} \approx 1839$$

Pour les atomes "légers" (comme H) et à plus forte raison pour les atomes "lourds" (comme Pb) la masse des électrons est toujours négligeable.

La masse de l'atome est concentrée dans son noyau.

Exercice 2 :

Cu : $Z = 29$

Isotope 1 : $M_1 = 62,929 \text{ g mol}^{-1} \Leftrightarrow A_1 = 63 \Rightarrow N_1 = 34$

29 protons ; 29 électrons et 34 neutrons

Isotope 2 : $M_2 = 64,927 \text{ g mol}^{-1} \Leftrightarrow A_2 = 65 \Rightarrow N_2 = 36$

29 protons ; 29 électrons et 36 neutrons

$$M = \sum x_i M_i$$

$$M_{\text{Cu}} = x_1 M_1 + x_2 M_2$$

$$\sum x_i = 1 \Rightarrow x_1 + x_2 = 1 \Rightarrow x_2 = 1 - x_1$$

$$M_{\text{Cu}} = x_1 M_1 + (1 - x_1) M_2 = x_1 M_1 + M_2 + x_1 M_2$$

$$M_{\text{Cu}} - M_2 = x_1 (M_1 - M_2)$$

$$x_1 = (M_{\text{Cu}} - M_2) / (M_1 - M_2) = (63,540 - 64,927) / (62,929 - 64,927) = 0,6942$$

Exercice Corrig Atomistique Examen Corrig

JR Anderson



Exercice Corrig Atomistique Examen Corrig:

Chimie physique Elisabeth Le Masne de Chermont, Arnaud Le Masne de Chermont, 1998 **Atomistique et liaison chimique** Yves Jean, François Volatron, 1995 Ce livre est destiné aux élèves de première année des Classes Préparatoires aux Grandes Ecoles filière scientifique Il recouvre la partie du nouveau programme consacrée à l'étude de la structure électronique des atomes et des molécules Conformément à l'esprit de ce programme la structure électronique des atomes et des molécules y est présentée d'une façon qualitative non mathématique qui privilégie les concepts de base et les raisonnements Toutes les notions utiles même les plus élémentaires sont introduites progressivement de façon à donner une cohérence à l'ouvrage et rendre sa lecture aisée L'ensemble est illustré de nombreux schémas et figures qui facilitent grandement la compréhension Enfin chaque chapitre est complété par un questionnaire Vrai ou Faux et par une série d'exercices soigneusement choisis qui permettent au lecteur de vérifier l'acquisition de ses connaissances et de les appliquer des problèmes nouveaux **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1972 **Livres hebdo**, 2001 **La chimie dans les grandes écoles et classes scientifiques préparatoires** Théophile Mbang, Thierry Stéphane Ndem Mbang, 2012-01-01 Cet ouvrage complet et très clair conçu pour guider l'étudiant dans son apprentissage de la chimie est illustré de nombreux exemples pour comprendre et assimiler Il s'adresse plus spécifiquement aux étudiants du 1er cycle des grandes écoles d'ingénieurs Ecoles normales supérieures d'enseignement général et technique facultés de génie industriel classes préparatoires **Les Livres disponibles**, 2004 La liste exhaustive des ouvrages disponibles publiés en langue française dans le monde La liste des éditeurs et la liste des collections de langue française Le Catalogue de l'édition française, 1976 Une liste exhaustive des ouvrages disponibles publiés en français de par le monde **Books and Pamphlets, Including Serials and Contributions to Periodicals** Library of Congress. Copyright Office, 1969 *La Recherche*, 1992 **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1969 Bulletin des sciences mathématiques, 1827 Bulletin général et universel des annonces et des nouvelles scientifiques Bulletin des sciences mathématiques, astronomiques, physiques et chimiques, 1827 Chimie générale Maria Gruia, Michèle Polisset, 2001 La collection La Chimie en poche vise à présenter par l'intermédiaire de rappels de cours de preuves et d'exercices corrigés et commentés par des enseignants chevronnés chacun des modules principaux de l'enseignement de Chimie en première année de médecine mais aussi en DEUG A et B et en première année de faculté de pharmacie Le lecteur y trouvera un véritable mode d'emploi de sa réussite aux examens et concours *Bulletin general et universel des annonces et des nouvelles scientifiques, publie sous la direction du baron de Ferussac* André-Étienne-Just-Pascal-Joseph-François D'Audebard "de" Férussac, 1827 Bulletin universel des sciences et de l'industrie, 1827 *Bulletin universel des sciences et de l'industrie. 1*, 1827 **Bulletin universel des sciences et de l'industrie** Camille Deflers, Bulos (M), Chevillot, Jacques Frédéric Saigey, Charles Sturm (i.e. Jacques Charles François), Henry François Gaultier de Claubry, 1827 *Atomistique* E. B. MAAROUF, 1999

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