

# 1. CONDUCTOMETRIC TITRATIONS

**Theory.** According to Kohlrausch, the electrical conductance depends upon the numbers and mobilities of ions. Let us consider the titration of a strong acid like HCl with a strong base like NaOH. Before the NaOH solution is added the acid solution has a high conductance due to the highly mobile hydrogen ions. As alkali solution is added the hydrogen ions are removed by combination with the hydroxyl ions forming feebly ionised water molecules and their place is taken by the much slower moving  $\text{Na}^+$  ions.  $[\text{H}^+ + \text{Cl}^-] + [\text{Na}^+ + \text{OH}^-] \rightarrow \text{Na}^+ + \text{Cl}^- + \text{H}_2\text{O}$  (feebly ionised) Consequently the conductance of the solution decreases and continues falling off with addition of solution hydroxide until the

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equivalence point is reached. Any subsequent addition of alkali means increase of  $\text{Na}^+$  ions and the fast moving hydroxyl ions and thus the conductance begins to increase. If we plot the conductance of the solution against the volume of alkali added the points will lie on two lines which are almost straight. The point of intersection of intrapolated lines [Fig. 12-8 (a)] will be the end point of titration and will determine the volume of alkali required for neutralisation of the acid. Compare the slopes of the two straight lines and state reasons for the difference in the slopes.

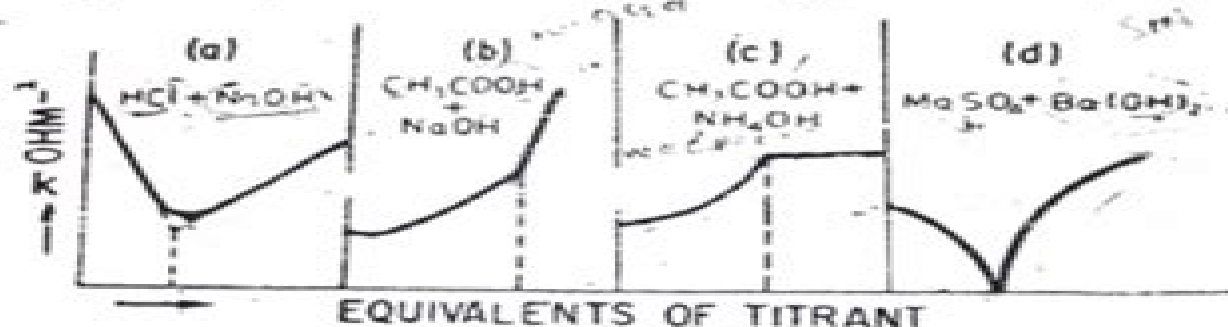


Fig. 12-8. Typical Conductometric Titration Curves.

If we titrate a weak acid like acetic acid with a strong base like sodium hydroxide we will get a curve like (b).

# Practical Reports On Conductometric Titrations

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**Chemical Engineering and Mining Review** ,1927     **Some New Developments in Agricultural Science** Eva May Davis,F. L. Duley,Frederick Blackmar Mumford,Henry Herman Krusekopf,J. T. Quinn,John Cochran Wooley,Leonard Dixon Haigh,Thomas Jesse Talbert,1925

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