

## Chapter 3 – Stoichiometry

**Question 3.1:** Determine the molecular/molar masses of the following molecules:

- |   |                               |
|---|-------------------------------|
| a. $\text{C}_2\text{H}_6$ 30.07 g/mol                         | d. $\text{NaOH}$ 39.99 g/mol  |
| b. $\text{Sulfuric acid (H}_2\text{SO}_4\text{)}$ 98.08 g/mol | e. $\text{HNO}_3$ 63.01 g/mol |
| c. $\text{C}_2\text{H}_5\text{OH}$ 46.07 g/mol                | f. $\text{CaO}$ 56.08 g/mol   |

**Question 3.2:** Determine the molecular/molar mass of the following molecules:

- |  |                                       |
|--|---------------------------------------|
| a. $\text{H}_2\text{SO}_4$ 98.08 g/mol | d. $\text{H}_2\text{O}$ 18.02 g/mol   |
| b. $\text{NaOH}$ 39.99 g/mol           | e. $\text{NaOH}$ 39.99 g/mol          |
| c. $\text{NaOH}$ 39.99 g/mol           | f. $\text{Sulfuric acid}$ 98.08 g/mol |

**Question 3.3:** Express the following grams of each component in moles of each component:

- |  |  |
|--|--|
| a. 27.0 kg $\text{C}_2\text{H}_6$ 1.0 kmol   | d. 1.0 kg $\text{Sulfuric acid}$ 0.0103 kmol |
| b. 1.0 kg $\text{C}_2\text{H}_6$ 0.0357 kmol | e. 10 kg $\text{H}_2\text{O}$ 0.555 kmol     |
| c. 10.0 kg $\text{CaO}$ 0.178 kmol           | f. 10 kg $\text{CaO}$ 0.178 kmol             |

**Question 3.4:** Express the following grams of each component in moles of each component:

- |   |  |
|---|--|
| a. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol | d. 10 kg $\text{Sulfuric acid}$ 0.103 kmol |
| b. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol | e. 10 kg $\text{Sulfuric acid}$ 0.103 kmol |
| c. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol | f. 10 kg $\text{Sulfuric acid}$ 0.103 kmol |

**Question 3.5:** Express the following grams of each component in moles of each component:

- |  |  |
|--|--|
| a. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol  | d. 1.0 kg $\text{Sulfuric acid}$ 0.0103 kmol |
| b. 1.0 kg $\text{C}_2\text{H}_6$ 0.0357 kmol | e. 10 kg $\text{H}_2\text{O}$ 0.555 kmol     |
| c. 10.0 kg $\text{CaO}$ 0.178 kmol           | f. 10 kg $\text{CaO}$ 0.178 kmol             |

**Question 3.6:** Express the following grams of each component in moles of each component:

- |   |   |
|---|---|
| a. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol | d. 10 kg $\text{Sulfuric acid}$ 0.103 kmol  |
| b. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol | e. 10 kg $\text{Sulfuric acid}$ 0.103 kmol  |
| c. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol | f. 10 kg $\text{H}_2\text{SO}_4$ 0.103 kmol |

**Question 3.7:** Balance the following chemical reactions:

- |  |  |
|--|--|
| a. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 1 kmol | d. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 1 kmol |
| b. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 1 kmol | e. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 1 kmol |
| c. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 1 kmol | f. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 1 kmol |

# Chapter 3 Stoichiometry Exercises Answers

**Steven S. Zumdahl**



### Chapter 3 Stoichiometry Exercises Answers:

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