

6. The new penny is lighter because it is mostly zinc which has a lower density than copper.

You're The Chemist

- at 90°, mass of 1 drop: 0.019 g
at 45°, mass of 1 drop: 0.0218 g
at 0°, mass of 1 drop: 0.0242 g
Pipets give different results.
- The best angle is 90° because the pipet is easiest to control. Expel the air bubble so that the first drop will be the same size as the rest.
- Find mass of can and divide by density of aluminum. Sample answer: mass of one can: 14.77 g;
 d_{Al} : 2.70 g/cm³; $V = 5.47 \text{ cm}^3$
- (1) Measure the mass before and after you fill the can with water. Use the mass and density of water to find the volume.
(2) Measure the height and radius and calculate volume.
 $V = \pi r^2 h$
(Can is not a perfect cylinder.)
(3) Read label: 12 oz = 355 mL.
- Sample answer:
 $V = 16.5 \text{ m} \times 3.0 \text{ m} \times 12.8 \text{ m}$
 $= 630 \text{ m}^3 \times 1000 \text{ L/m}^3$
 $= 630,000 \text{ L}$
Assume 30 people with an average weight of 130 lb (1 kg = 2.2 lb) and a density of about 1.0 kg/L.
Volume of 30 people:
 $V = 30 \times 130 \text{ lb} \times 1 \text{ kg}/2.2 \text{ lb} \times 1 \text{ L}/1.0 \text{ kg}$
 $= 1800 \text{ L}$
The volume of 30 chairs, 15 tables, and 2 desks is about that of 30 people or 1800 L.
The volume of people and furniture is 3600 L.
 $\% \text{ error} = (3600 \text{ L}/630,000 \text{ L})(100\%)$
 $= 0.57\%$.
- If die measures 1.55 cm on a side:
 $V = (1.55 \text{ cm})^3 = 3.72 \text{ cm}^3$
A die has 21 holes that are hemispheres with a radius of 0.20 cm.
Volume of hemisphere: $2/3\pi r^3$
 $V = 2/3\pi r^3 = 0.017 \text{ cm}^3$
Volume of 21 hemispheres: 0.36 cm^3
Volume of die:

$$V = 3.72 \text{ cm}^3 - 0.36 \text{ cm}^3$$

$$= 3.36 \text{ cm}^3$$

$$\text{Error} = 0.36 \text{ cm}^3$$

Percent error

$$= (0.36 \text{ cm}^3/3.36 \text{ cm}^3)(100\%) = 11\%$$

Note: The holes in some dice are cones. The volume of a cone is $1/3 \pi r^2 h$.

- Find weight in pounds and convert to kg.
Assume the density is about 1.00 kg/L.
 $V = \text{weight} \times 1 \text{ kg}/2.2 \text{ lb} \times 1 \text{ L}/1.00 \text{ kg}$

Section Review 4.1

Part A Completion

- | | |
|--------------|-------------------------------------|
| 1. atoms | 3. compounds |
| 2. different | 4. separated, joined, or rearranged |

Part B True-False

- | | | |
|-------|-------|-------|
| 5. NT | 6. ST | 7. AT |
|-------|-------|-------|

Part C Matching

- | | |
|------|-------|
| 8. c | 10. d |
| 9. a | 11. b |

Part D Questions and Problems

- simple whole-number ratios
- $\frac{1 \times 10^6 \text{ atoms}}{\text{cm}^3} \times \frac{100 \text{ cm}^3}{\text{m}} = 10^{10} \text{ atoms/m}$

Section Review 4.2

Part A Completion

- | | |
|--------------|----------------------|
| 1. subatomic | 5. nucleus |
| 2. negative | 6. Ernest Rutherford |
| 3. protons | 7. positive |
| 4. neutrons | 8. electrons |

Part B True-False

- | | |
|--------|--------|
| 9. NT | 11. NT |
| 10. AT | 12. AT |

Part C Matching

- | | | |
|-------|-------|-------|
| 13. c | 15. e | 17. b |
| 14. a | 16. d | |

Part D Questions and Problems

- protons, neutrons

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Table of Contents Chapter 4 Assessment Chemistry Answers Pearson

1. Understanding the eBook Chapter 4 Assessment Chemistry Answers Pearson
 - The Rise of Digital Reading Chapter 4 Assessment Chemistry Answers Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Assessment Chemistry Answers Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Assessment Chemistry Answers Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Assessment Chemistry Answers Pearson
 - Personalized Recommendations
 - Chapter 4 Assessment Chemistry Answers Pearson User Reviews and Ratings
 - Chapter 4 Assessment Chemistry Answers Pearson and Bestseller Lists

5. Accessing Chapter 4 Assessment Chemistry Answers Pearson Free and Paid eBooks
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 - Chapter 4 Assessment Chemistry Answers Pearson eBook Subscription Services
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 - Chapter 4 Assessment Chemistry Answers Pearson Enhanced eBook Features
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 - Highlighting and Note-Taking Chapter 4 Assessment Chemistry Answers Pearson
 - Interactive Elements Chapter 4 Assessment Chemistry Answers Pearson
8. Staying Engaged with Chapter 4 Assessment Chemistry Answers Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Assessment Chemistry Answers Pearson
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Assessment Chemistry Answers Pearson
 - Setting Reading Goals Chapter 4 Assessment Chemistry Answers Pearson
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Assessment Chemistry Answers Pearson
 - Fact-Checking eBook Content of Chapter 4 Assessment Chemistry Answers Pearson
 - Distinguishing Credible Sources
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- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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