

## Chapter 4 Atomic Structure

**Section 4.2 The Structure of an Atom****(pages 108–112)**

*This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.*

**Reading Strategy (page 108)**

**Monitoring Your Understanding** Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

**Properties of Subatomic Particles (pages 108–109)**

- What are three subatomic particles?  
a. Protons      b. Electrons      c. Neutrons
- Circle the letter that identifies a subatomic particle with a positive charge.  
a. nucleus      **(b)** proton  
c. neutron      d. electron
- Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.  
\_\_\_\_\_  
\_\_\_\_\_

**Comparing Subatomic Particles (pages 109–110)**

- Circle the letters of properties that vary among subatomic particles.  
a. color      **(b)** mass  
**(c)** charge      **(d)** location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.  
**(a)** mass of 1 neutron = mass of 1 proton  
b. mass of 2000 neutrons = mass of 1 proton  
c. mass of 1 electron = mass of 1 proton  
d. mass of 1 neutron = mass of 1 electron

# Chapter 4 Atomic Structure Worksheet Answers

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six domains Concepts Processes Applications Attitudes Creativity and World view of Science to make teaching and learning of Chemistry interesting understandable and enjoyable for young minds This series builds a solid foundation for young learners to prepare them for higher classes The main strength of the series lies in the subject matter and the experience that a learner will get in solving difficult and complex problems of Chemistry Emphasis has been laid upon mastering the fundamental principles of Chemistry rather than specific procedures Unique features of this series are The content of the book is written in a very simple and easy to understand language All the Key concepts in the curriculum have been systematically covered and graded in the text Each theme has been divided into units followed by thought provoking and engaging exercises to test the knowledge understanding and applications of the concepts learnt in that unit At the end of each theme a comprehensive theme assignment which is aligned with the guidelines provided in National Education Policy NEP 2020 is given Explanations illustrations diagrams experiments and solutions to numerical problems have been included to make the subject more interesting comprehensive and appealing Diagrams illustrations and text have been integrated to enhance comprehension Definitions and other important scientific information are highlighted Throughout the series investigations related to the text enable the learners to learn through experimentation Quick revision of each chapter has been given under the caption Highlights in Review Online Support It provides Video lectures Unit wise interactive exercises Chapterwise Worksheet Solution of textbook questions for Teachers only E Book for Teachers only I hope this series would meet the needs and requirements of the curriculum to achieve the learning outcomes as laid down in the curriculum Suggestions and constructive feedback for the further improvement of the book shall be gratefully acknowledged and incorporated in the future edition of the book Author      **Learning Elementary Chemistry for Class 7 (A.Y.**

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