

Chapter 9 Practice Problems: Chemical Names and Formulas

Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
 - oxygen
 - iodine
 - sodium
 - aluminum
 - nickel, two electrons lost
 - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
 - Cr^{3+}
 - P^{3-}
 - Li^+
 - Ca^{2+}
 - Cl^-
 - O^{2-}
- Name each ion. Identify each as a cation or an anion.
 - Sn^{2+}
 - Co^{3+}
 - Br^-
 - K^+
 - H^-
 - Mn^{2+}
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
 - carbonate ion
 - nitrite ion
 - sulfate ion
 - hydroxide ion
 - chromate ion
 - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
 - CN^-
 - HCO_3^-
 - PO_4^{3-}
 - Cl^-
 - Ca^{2+}
 - SO_3^{2-}

9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
 - magnesium oxide
 - tin(II) fluoride
 - potassium iodide
 - aluminum chloride
 - sodium sulfide
 - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
 - Ba^{2+} and Cl^-
 - Ag^+ and I^-
 - Ca^{2+} and S^{2-}
 - K^+ and Br^-
 - Al^{3+} and O^{2-}
 - Fe^{2+} and O^{2-}
- Name the following binary ionic compounds.
 - MnO_2
 - Li_3N
 - CaCl_2
 - SrBr_2
 - NiCl_2
 - K_2S
 - Fe_3N_2
 - SnCl_4
- Write formulas for the following ionic compounds.
 - sodium phosphate
 - magnesium sulfate
 - sodium hydroxide
 - potassium cyanide
 - ammonium chloride
 - potassium dichromate
- Write formulas for the following ionic compounds.
 - NH_4^+ and SO_4^{2-}
 - K^+ and NO_3^-
 - barium ion and hydroxide ion
 - lithium ion and carbonate ion

Chapter 9 Chemical Names Formulas Practice Problems Answers

Fred M. Dewey



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General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features Understanding Chemistry Fred M. Dewey, 1994 **Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed** Peter P. Berlow, 1978 **CliffsNotes Chemistry Practice Pack** Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level *Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book **eBook: General, Organic and Biological Chemistry 2e** SMITH, 2012-02-16 eBook General Organic and Biological Chemistry 2e Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Addison-Wesley Chemistry**, 1990 The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB **Partial Solutions Manual** Darrell D. Ebbing, David Bookin, 1998 **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and

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Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students

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