

Chapter 21: Nuclear Chemistry21.1: The Nature of Nuclear Reactions

Nucleons: - the particles that make up a nucleus of an atom (protons, (${}^1_1p^+$ or 1_1H) and neutrons, (1_0n)).

Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

Nuclide: a particular atom or isotope containing specific numbers of protons and neutrons



Radioactivity: - the particles and/or electromagnetic radiation that are emitted due to unstable nuclei.
 - all elements having atomic number 84 (Polonium) and greater are radioactive.

Nuclear Transmutation: - a process where radioactivity is resulted from the bombardment of nuclei by neutrons, protons or other nuclei.
 - in most cases, heavier elements are synthesized from lighter elements.

History of Radioactivity

- In 1896, **Wilhelm Roentgen** discovered X-ray by examining ray emitting from the outside of the cathode ray glass tube. It has the capability of passing through solid materials, but can be blocked by denser matters. It can also be exposed to photographic plate, resulting from an image of "seeing through" an container made of less dense material.
- In the same year, **Antoine Becquerel** discovered that uranium emits a ray onto a photographic plate in the absence of sunlight or other forms of energy.
- A few years later, **Marie and Pierre Curie** demonstrated that radiation can be emitted by other elements. They discovered two new elements (polonium and radium) based on their tendency to emit radiation (**radioactivity**).

Types of Radioactive Particle and Decay

- Alpha Particle (α particle):** - basically a helium nucleus (4_2He), commonly found during radioactive decay from heavier nuclide (the net result is to increase the neutron to proton ratio – more explanation in the next section).



- Beta Particle (β particle):** - basically an electron (${}^0_{-1}e$ or ${}^0_{-1}\beta$) that is emitted when the neutron to proton ratio is higher than the zone of stability (**a neutron is transformed to a proton and an electron as a result** – more explanation in the next section).
 - electrons have a mass number of 0 and an atomic number assignment of -1, due to its charge.



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **A Complete Introduction to Modern NMR Spectroscopy** Roger S. Macomber, 1997-12-23 Clear accessible coverage of modern NMR spectroscopy for students and professionals in many fields of science Nuclear magnetic resonance NMR spectroscopy has made quantum leaps in the last decade becoming a staple tool in such divergent fields as chemistry physics materials science biology and medicine That is why it is essential that scientists working in these areas be fully conversant with current NMR theory and practice This down to basics text offers a comprehensive up to date treatment of the fundamentals of NMR spectroscopy Using a straightforward approach that develops all concepts from a rudimentary level without using heavy mathematics it gives readers the knowledge they need to solve any molecular structure problem from a complete set of NMR data Topics are illustrated throughout with hundreds of figures and actual spectra Chapter end summaries and review problems with answers are included to help reinforce and test understanding of key material From NMR studies of biologically important molecules to magnetic resonance imaging this book serves as an excellent all around primer on NMR spectroscopic analysis

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Table of Contents Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern

1. Understanding the eBook Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - The Rise of Digital Reading Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Personalized Recommendations

- Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern User Reviews and Ratings
- Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern and Bestseller Lists
- 5. Accessing Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern Free and Paid eBooks
 - Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern Public Domain eBooks
 - Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern eBook Subscription Services
 - Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern Budget-Friendly Options
- 6. Navigating Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern Compatibility with Devices
 - Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Highlighting and Note-Taking Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Interactive Elements Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
- 8. Staying Engaged with Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
- 9. Balancing eBooks and Physical Books Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Setting Reading Goals Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern
 - Fact-Checking eBook Content of Chapter 22 Review Nuclear Chemistry Section 2 Answers Modern

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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