Chemical Bonding

MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

bond energy b. covalent bond enthalpy of vaporization c. metallic bond

Use the electronegativity values shown in Figure 20, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

 ionic
 a. H—F
 nonpolar covalent
 d. H—H

 ionic
 b. Na—Cl
 polar covalent
 e. H—C

 polar covalent
 c. H—O
 polar covalent
 f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.
Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

Chapter 8 Covalent Bonding Assessment Answers

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Chapter 8 Covalent Bonding Assessment Answers:

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Adopting the Tune of Term: An Psychological Symphony within Chapter 8 Covalent Bonding Assessment Answers

In a global consumed by monitors and the ceaseless chatter of immediate transmission, the melodic splendor and emotional symphony developed by the prepared term usually disappear in to the backdrop, eclipsed by the persistent noise and interruptions that permeate our lives. Nevertheless, situated within the pages of **Chapter 8 Covalent Bonding Assessment Answers** a charming literary value full of natural emotions, lies an immersive symphony waiting to be embraced. Crafted by an elegant musician of language, this charming masterpiece conducts viewers on an emotional journey, skillfully unraveling the concealed melodies and profound influence resonating within each carefully crafted phrase. Within the depths of the emotional review, we will investigate the book is central harmonies, analyze their enthralling publishing model, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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