

### Types of Chemical Reactions Worksheet #1

#### **11. Chemistry**

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  **Synthesis**
  - 2)  $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$  **Single Replacement**
  - 3)  $\text{HNO}_3 + \text{AlCl}_3 \rightarrow \text{Al}(\text{OH})_3 + 3\text{HCl}$  **Double Replacement**
  - 4)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  **Synthesis**
  - \*5)  $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  **Double Replacement**
  - 6)  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$  **Decomposition**
  - 7)  $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$  **Single Replacement**
  - 8)  $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$  **Double Replacement**
  - 9)  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$  **Synthesis**
  - \*10)  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  **Single Replacement**
- \*Remember: "H<sub>2</sub>O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 1.1) Double replacement:  $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 1.2) Single replacement:  $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 1.3) Decomposition:  $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 1.4) Synthesis:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 1.5) Double replacement:  $3\text{Ca}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CaCl}_2$
- 1.6) Single replacement:  $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$  [Assume Fe<sup>2+</sup>]
- 1.7) Decomposition:  $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 1.8) Synthesis:  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 1.9) Double replacement:  $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 2.0) Single replacement:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$  [Copper (II) is used here]

# Chemfax Lab Answers Chemical Reactions

**John Joseph Griffin**



## Chemfax Lab Answers Chemical Reactions:

**Fizz, Gurgle, Pop! Chemical Reactions in the Lab** Daniel R. Faust, 2020 Whether they are testing a new drug creating new materials or checking the safety of our food and water all scientists follow the scientific method One of the major steps in the scientific method is testing your hypothesis That s when scientists hit the lab This volume stresses the important of safety in the laboratory whether it s a school science lab or the most high tech research facility It also examines the proper rules and procedures everyone must follow when working with chemicals in the lab Fact boxes and full color photographs support the clear age appropriate text *Chemical Reactions and Their Equations* Ingo Waldemar Dagobert Hackh, 1928

*Chemical Reactions Science Learning Guide* NewPath Learning, 2014-03-01 The Chemical Reactions Student Learning Guide includes self directed readings easy to follow illustrated explanations guiding questions inquiry based activities a lab investigation key vocabulary review and assessment review questions along with a post test It covers the following standards aligned concepts Changes of Matter Chemical Reactions Formulas Balancing Equations Types of Chemical Reactions 1 Types of Chemical Reactions 2 Energy in Chemical Reactions Evidence of Chemical Reactions and Chemical Reaction Rates Catalysts Aligned to Next Generation Science Standards NGSS and other state standards Chemical reactions and their equations Ingo Waldemar Dagobert Hackh, 1921 **Chemical Reactions** Kristi Lew, 2009 Explores how chemical reactions happen from how your body breaks down the food you eat to how forensic experts use them to solve crimes *CHEMICAL REACTIONS AND THEIR EQUATIONS* INGO W. D. HACKH, 2018 *Structure and Change* Alan C. Blizzard, Richardson, R. A. (Robert A.), Humphreys, D. A. (David A.), 1981 Chemical Reactions and Their Equations Ingo Waldemar Dagobert Hackh, 2016-05-17 This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work was reproduced from the original artifact and remains as true to the original work as possible Therefore you will see the original copyright references library stamps as most of these works have been housed in our most important libraries around the world and other notations in the work This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work As a reproduction of a historical artifact this work may contain missing or blurred pages poor pictures errant marks etc Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant *Chemical Reactions* John A. Olson, 2011-06-08 **Chemical Reactions** Denise Walker, 2007 This title introduces the reader to the huge variety of chemical reactions that shape our world Find out all about explosions learn about how to start reactions and understand how chemical equations work **Chemical Reactions** John Joseph Griffin, 1854 Chemical Reactions John Joseph Griffin, 1849 **Chemical Reactions** John Joseph Griffin, 2019 Chemical Reactions Mr. Ajeet Singh, 2018-05-22 A

fundamental text covering types of chemical reactions reaction energetics mechanisms and industrial applications

Practice Makes Perfect Chemistry Review and Workbook, Second Edition Marian DeWane, Heather Hattori, 2018-12-28

The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you ve learned along with a complete answer key to check your work Throughout this book you will learn the terms to help you understand chemistry and you will expand your knowledge of the subject through hundreds of sample questions and their solutions With the lessons in this book you will find it easier than ever to grasp chemistry concepts And with a variety of exercises for practice you will gain confidence using your growing chemistry skills in your classwork and on exams YOU LL BE ON YOUR WAY TO MASTERING THESE TOPICS AND MORE Atomic structure The periodic table Chemical formulas Chemical reactions Mass and mole relationships Gas laws Solutions Acids and bases Thermochemistry A brand new chapter on the structure of molecules      **Chemical Reactions** Rachael Morlock, 2020      **Bridges: Chemistry in the Kitchen ,      Exploring Chemical Reactions** Nigel Saunders, 2007

Chemical Reactions Anne O'Daly, 2026 This book looks at chemical reactions the processes in which the atoms of substances rearrange to make different substances      **Encyclopedia of Chemical Reactions** C. A. Jacobson, 1946

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