AP Chemistry

Dr. Mandes

Name:
Period:
Date:

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

- Please do #18 in chapter 12 of your text.
 - a. This increases the [H₂], which will increase the rate, but has no effect on k
 - b. Due to Ambenius equation, changing temperature changes the value of k.
 - Catalyst function by lowering the activation energy, so due to Amhenius equation, changing the activation energy changes the value of k.
- Consider the reaction: P₄ + 6 H₂ → 4 PH₅. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

{P ₄ }, mol/L	[H ₂], mol/L	Initial Rate, mol/(L 's)
0.0110	0.0075	3.20 × 10 ⁻⁴
0.0110	0.0150	6.40 × 10 ⁻⁴
0.0220	0.0150	6.39 × 10 ⁻⁴

- a. What is the order with respect to: $P_4 = 0$.
 - H₂ _______
- b. Write the rate law for this reaction. rate = k[H₂]
- c. Determine the value and units of the rate constant, k. plug and chug using the rate law & data from exp't 1 and solving for k, we get k = 0.0427 s'
- Consider the reaction: SO₂ + O₃ → SO₃ + O₂. A rate study of this reaction was conducted at 298 K. The
 data that were obtained are shown in the table.

[SO ₂], mol/L	[O ₅], mol/L	Initial Rate, mol/(L 's)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1,062

- b. Write the rate law for this reaction. rate = k[SO₂]²[O₃]⁶
- e. Determine the value and units of the rate constant, k. plug and chug using the rate law & data from exp"t 1 and solving for k, we get k = 2.36 mol L.** s.*
- 8. Consider the following mechanism. $A_2 + B_3 \rightarrow \mathbb{R} + \mathbb{C}$ (slow)

$$A_2 + R \rightarrow C$$
 (fast)

- a. Write the overall balanced chemical equation.
 2 A₂
 B₂
 2 C
- b. Identify any intermediates within the mechanism. R
- c. What is the order with respect to each reactant? A21"; B2 1"

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