

Equilibrium Worksheet **SOLUTIONS**

Complete the following questions on a separate piece of paper.

1. Write the equilibrium expression, K_{eq} , for each of the following reactions:



$$K_{eq} = \frac{[\text{NO}_2]^2}{[\text{NO}]^2[\text{O}_2]}$$



$$K_{eq} = \frac{[\text{H}_2\text{O}]^2[\text{Cl}_2]^2}{[\text{HCl}]^4[\text{O}_2]}$$



$$K_{eq} = \frac{[\text{NO}][\text{Cl}_2]^{\frac{1}{2}}}{[\text{NOCl}]}$$



$$K_{eq} = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$$



$$K_{eq} = [\text{Ca}^{2+}][\text{Cl}^{-}]^2$$



$$K_{eq} = \frac{[\text{H}_3\text{O}^{+}][\text{C}_2\text{H}_3\text{O}_2^{-}]}{[\text{HC}_2\text{H}_3\text{O}_2]}$$



$$K_{eq} = \frac{1}{[\text{CO}_2]}$$



$$K_{eq} = \frac{[\text{COCl}_2]}{[\text{CO}_2][\text{Cl}_2]^2}$$



$$K_{eq} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2$$



$$K_{eq} = \frac{[\text{CO}_2]}{[\text{CO}]}$$

2. The equilibrium constant for the equilibrium $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{H}_{2(g)}$ is 302 at 600K. What is the value of the equilibrium constant for the reverse reaction at the same temperature?

$$K_{eq, \text{reverse}} = \frac{1}{K_{eq}} = \frac{1}{302} = 3.31 \times 10^{-3}$$

3. Classify the following equilibria as heterogeneous or homogeneous, and write an equilibrium expression for each.



heterogeneous

$$K_{eq} = [\text{N}_2][\text{H}_2\text{O}]^2$$

Chemical Equilibrium Worksheet With Answers

Wendy Ward Hoffer



Chemical Equilibrium Worksheet With Answers:

General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

Illinois Chemistry Teacher, 1992

Science as Thinking Wendy Ward Hoffer, 2009 You are about to immerse yourself in a gorgeously readable and engaging account of how teachers can move science instruction from hands on to minds on Wendy Ward Hoffer describes how teachers can extrapolate what is known about good thinking strategies instruction to science teaching and learning Hoffer illuminates the path for thousands of teachers in science and beyond who today work with those who will lead this country's efforts in energy health care the exploration of sea and space and the protection of our planet What work is more vital to our future

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The Science Teacher, 2008 SCC Library has 1964 cur

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Concepts of Biogeography &

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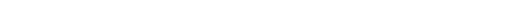
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