

## Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

### District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	<sup>184</sup> W <sup>+2</sup> 74	74	74	110	72	184	+2

- Name the compound  
 a.  $\text{Fe}_2\text{O}_3$  – **iron (III) oxide**  
 b.  $\text{Cl}_2\text{O}_8$  – **dichlorine octaoxide**
- Write the formula  
 a. Trisulfur Dinitride -  **$\text{S}_3\text{N}_2$**   
 b. Zinc Hydroxide –  **$\text{Zn}(\text{OH})_2$**   
 c. Tin (IV) oxide -  **$\text{SnO}_2$**
- The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? Volume is constant.  

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{843\text{mmHg}}{297\text{K}} = \frac{P_2}{273\text{K}} \quad P_2 = 775 \text{ mm Hg}$$
- The volume of a gas is 81.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?  

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \quad \frac{(119\text{kPa})(81.7\text{mL})}{346\text{K}} = \frac{(101.3\text{kPa})(V_2)}{273\text{K}} \quad V_2 = 75.7 \text{ mL}$$
- Between Be and Ba, which has a larger ionization energy? Larger atomic radius?  
**Larger ionization energy – Be**  
**Larger atomic radius - Ba**
- Balance the equation & classify the type of reaction:  

$$2\text{Al} + 3\text{CuSO}_4 \rightarrow 1\text{Al}_2(\text{SO}_4)_3 + 3\text{Cu} \quad \text{single replacement}$$
- Predict the products & classify the type of reaction:  

$$\text{H}_3\text{PO}_4 + 3\text{NaOH} \rightarrow 3\text{H}_2\text{O} + \text{Na}_3\text{PO}_4 \quad \text{double replacement}$$
- Calculate the Molar Mass:  $(\text{NH}_4)_2\text{CO}_3$   
**96 g/mole**
- Convert 2.75 moles of  $\text{N}_2\text{O}_5$  to molecules  

$$\frac{2.75\text{moles}}{1} \times \frac{6.02 \times 10^{23}\text{molecules}}{1\text{mole}} = 1.66 \times 10^{24} \text{ molecules } \text{N}_2\text{O}_5$$

### Solutions

# Chemistry Spring Semester Final Exam Review Sheet

**Ruskin Skidmore Freer**



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**Survival Handbook for the New Chemistry Instructor** Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

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