



- a. combination  
b. decomposition

- c. displacement  
d. exchange

19. Classify the following reaction.



- a. combination  
b. decomposition

- c. displacement  
d. exchange

20. The Roman numerals in the reaction given represent the coefficients in the balanced chemical equation. What are the values of the coefficients?



- I II III IV  
a. 2 2 2 3  
b. 1 1 1 2  
c. 2 1 2 1  
d. 1 2 2 1  
e. 2 2 2 1

21. Which of the following is a diprotic acid?

- a.  $\text{H}_3\text{PO}_4$   
b.  $\text{H}_2\text{O}$   
c.  $\text{Ca(OH)}_2$

- d.  $\text{H}_2\text{SO}_4$   
e.  $\text{NH}_3$

22. Which compound can be classified as insoluble:

- a.  $\text{Mg(ClO}_4\text{)}_2$ ;  
b.  $\text{Cu(NO}_3\text{)}_2$ ;  
c.  $(\text{NH}_4)_2\text{SO}_4$ ;

- d.  $\text{Fe(OH)}_3$   
e.  $\text{NiSO}_4$ ;

23. Consider the reaction between 15.0 mL of a 1.00 M aqueous solution of  $\text{AgNO}_3$  and 10.0 mL of a 1.00 M aqueous solution of  $\text{K}_2\text{CO}_3$ . When these react, a precipitate is observed. What is present in solution after the reaction is complete (*Note: the precipitate is not considered to be in solution*)?

- a.  $\text{Ag}^+$ ,  $\text{NO}_3^-$ ,  $\text{K}^+$ ,  $\text{CO}_3^{2-}$ , water  
b.  $\text{Ag}^+$ ,  $\text{NO}_3^-$ ,  $\text{K}^+$ , water  
c.  $\text{K}^+$ ,  $\text{CO}_3^{2-}$ , water

- d.  $\text{NO}_3^-$ ,  $\text{K}^+$ ,  $\text{CO}_3^{2-}$ , water  
e. water

24. Which statement about strong acids is true?

- a. Strong acids are weak electrolytes.  
b. Strong acids are very concentrated.  
c. Strong acids are almost entirely converted to ions when dissolved in water.  
d. Acetic acid is a strong acid.  
e. All of the above are true.

25. Which of the following is a weak electrolyte?

- a.  $\text{HBr}$   
b.  $\text{Mg(OH)}_2$   
c.  $\text{NaCl}$

- d.  $\text{HCN}$   
e.  $\text{KI}$

26. What is the correct formula for the ammonium ion?

- a.  $\text{H}_2\text{O}$ ;  
b.  $\text{H}_3\text{O}^+$ ;  
c.  $\text{OH}^-$ ;

- d.  $\text{H}^+$ ;  
e.  $\text{NH}_4^+$ ;

27. What is the oxidation state (number) of Cl in  $\text{Cl}_2\text{O}_5$ ?

- a. +2  
b. -2

- c. -3  
d. -5

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**Michael Seilmaier**



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