

Key Equations for Colligative Properties

Raoult's Law: Relationship between the Vapor Pressure of a Solution (P_{solution}), the Mole Fraction of the Solvent (χ_{solvent}), and the Vapor Pressure of the Pure Solvent ($P^{\circ}_{\text{solvent}}$)

$$P_{\text{solution}} = \chi_{\text{solvent}} P^{\circ}_{\text{solvent}}$$

$$\chi_{\text{solvent}} = \frac{n_{\text{solvent}}}{n_{\text{solvent}} + n_{\text{solute}}}$$

The Vapor Pressure of a Solution Containing Two Volatile Components

$$P_A = \chi_A P^{\circ}_A \quad P_B = \chi_B P^{\circ}_B$$

$$P_{\text{total}} = P_A + P_B$$

Relationship between Osmotic Pressure (Π), Molarity (M), the Ideal Gas Constant (R), and Temperature (T , in K).

$$\Pi = MRT$$

$$M = \frac{n_{\text{solute}}}{V_{\text{solution}}} \quad 0.08206 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

The dependance of ΔT_f and ΔT_b on the Molality (m), and Freezing Point Depression Constant (K_f) and Boiling Point Elevation Constant (K_b) respectively.

$$\Delta T_f = m \cdot K_f$$

$$\Delta T_b = m \cdot K_b$$

$$m = \frac{\text{mol (solute)}}{\text{kg (solvent)}}$$

Colligative Properties Problems And Solutions

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Math Review Tool Kit Chem. Gary Leighton Long, Theodore L. Brown, Sharon D. Long, Doris Ingram Lewis, 1997

Chemical Calculations Sidney William Benson, 1971 Measurement Methods of measuring quantities of matter Chemical formulae Chemical reactions Energy and chemical changes The properties of gases The concept of combining power valence Measurement of solutions The physical properties of solutions Chemical equilibrium Electrolysis of ionic solutions Simple equilibria in ionic solutions The ionization of water hydrolysis Additional equilibria in ionic solutions Oxidation and reduction Predicting redox reactions Rates of chemical reactions **Math Review Toolkit** Gary L. Long, Michael P. O'Donnell, 2003-06

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