

Unit 2: Atoms and Elements  
AP Chemistry  
The Nuclear Atom

ALL of the answers to this worksheet can be **logically** figured out by looking at the **Schematic Diagrams for Various Atoms**, the **Periodic Table**, and **discussing** with your partners. All of the information you need is here somewhere. Determine each answer and be able to give convincing reasons for each answer. Good luck.

- How many **protons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **6**       $^{14}\text{C}$ ? **6**
- How many **neutrons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **7**       $^{14}\text{C}$ ? **7**
- How many **electrons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **6**       $^{14}\text{C}$ ? **7**
- Based on the model,
  - what do all carbon atoms (and ions) have in common?  
**Same number and types of protons.**
  - what do all hydrogen atoms (and ions) have in common?  
**Same number and types of protons.**
- What is the significance of the atomic number,  $Z$ , above each atomic symbol in the periodic chart?  
**It tells us the number of protons located in the nucleus of an atom of that particular element.**
- What do all nickel ( $\text{Ni}$ ) atoms have in common?  
**Same number and types of protons.**
- How is the mass number,  $A$ , (left-hand superscript next to the atomic symbol) determined?  
**By adding the protons and neutrons of an atom together.**
- What structural feature is different in isotopes of a particular element?  
**The number of neutrons that are present.**
- What feature distinguishes a neutral atom from an ion?  
**The number of electrons that are present.**
  - How is the charge on an ion determined?  
**By subtracting the number of electrons from the number of protons in the atom.**
- Where is most of the mass of an atom, within the nucleus or outside of the nucleus? Explain your reasoning.  
**Within the nucleus since that is where the protons and neutrons are located.**

11. Complete the following table:

Isotope	Atomic Number $Z$	Mass Number $A$	Number of electrons
$^{31}\text{P}$	15	31	15
$^{18}\text{O}$	8	18	8
$^{39}\text{Y}^{+}$	19	39	18
$^{58}\text{Ni}^{2+}$	28	58	26

# Chapter 22 Review Nuclear Chemistry Answer Key

**S Baum**



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