

## Chapter 6 Thermochemistry

### Homework Problems

- The enthalpy change for the fermentation (without oxygen,  $O_2$  gas) reaction of glucose to alcohol and carbon dioxide is  $-67 \text{ kJ}$ .
  - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
  - Is the reaction exothermic or endothermic?
  - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is  $0.71 \text{ J/}^\circ\text{C}\cdot\text{g}$ . Calculate the energy required to:
  - Raise the temperature of  $1.0 \text{ mol}$  graphite by  $1.0^\circ\text{C}$ .
  - Raise the temperature of  $850 \text{ g}$  graphite by  $150^\circ\text{C}$ .
  - Raise the temperature of  $75 \text{ kg}$  graphite from  $294 \text{ K}$  to  $348 \text{ K}$ .
- In a coffee-cup calorimeter,  $100.0 \text{ mL}$  of  $1.0 \text{ M NaOH}$  and  $100.0 \text{ mL}$  of  $1.0 \text{ M HCl}$  are mixed. Both solutions were originally at  $24.6^\circ\text{C}$ . After the reaction, the final temperature is  $31.3^\circ\text{C}$ . Assume that all the solutions have a density of  $1.0 \text{ g/mL}$  and a specific heat capacity of  $4.18 \text{ J/g}^\circ\text{C}$ . Calculate the enthalpy change for the neutralization of  $\text{HCl}$  by  $\text{NaOH}$ . Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ( $\text{C}_{15}\text{H}_{26}\text{O}$ ) has an energy of combustion of  $-5903.6 \text{ kJ/mol}$ . A sample of camphor with a mass of  $0.1204 \text{ g}$  is burned in a bomb calorimeter, the temperature increases by  $2.28^\circ\text{C}$ . Calculate the heat capacity of the calorimeter.

A  $0.1964 \text{ g}$  sample of quinone ( $\text{C}_6\text{H}_4\text{O}_2$ ) is burned in a bomb calorimeter that has a heat capacity of  $1.56 \text{ kJ/}^\circ\text{C}$ . The temperature of the calorimeter increases by  $3.2^\circ\text{C}$ . Calculate the energy of combustion of quinone per gram and per mole.

# Chapter 6 Thermochemistry Homework Problems

**H Kauffman**



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