Chemical Bonding

MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

bond energy b. covalent bond enthalpy of vaporization c. metallic bond

Use the electronegativity values shown in Figure 20, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

 ionic
 a. H—F
 nonpolar covalent
 d. H—H

 ionic
 b. Na—Cl
 polar covalent
 e. H—C

 polar covalent
 c. H—O
 polar covalent
 f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.
Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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G Psacharopoulos

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Table of Contents Chapter 8 Covalent Bonding Test A Answers Hazwoperore

- 1. Understanding the eBook Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - The Rise of Digital Reading Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Test A Answers Hazwoperore

- Personalized Recommendations
- Chapter 8 Covalent Bonding Test A Answers Hazwoperore User Reviews and Ratings
- Chapter 8 Covalent Bonding Test A Answers Hazwoperore and Bestseller Lists
- 5. Accessing Chapter 8 Covalent Bonding Test A Answers Hazwoperore Free and Paid eBooks
 - Chapter 8 Covalent Bonding Test A Answers Hazwoperore Public Domain eBooks
 - Chapter 8 Covalent Bonding Test A Answers Hazwoperore eBook Subscription Services
 - Chapter 8 Covalent Bonding Test A Answers Hazwoperore Budget-Friendly Options
- 6. Navigating Chapter 8 Covalent Bonding Test A Answers Hazwoperore eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chapter 8 Covalent Bonding Test A Answers Hazwoperore Compatibility with Devices
 - Chapter 8 Covalent Bonding Test A Answers Hazwoperore Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Highlighting and Note-Taking Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Interactive Elements Chapter 8 Covalent Bonding Test A Answers Hazwoperore
- 8. Staying Engaged with Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Covalent Bonding Test A Answers Hazwoperore
- 9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Test A Answers Hazwoperore
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Setting Reading Goals Chapter 8 Covalent Bonding Test A Answers Hazwoperore
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Test A Answers Hazwoperore

- Fact-Checking eBook Content of Chapter 8 Covalent Bonding Test A Answers Hazwoperore
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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