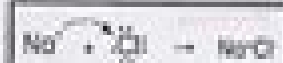


## IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1.  $\text{K} + \text{F}$



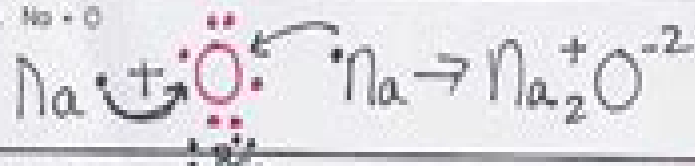
2.  $\text{Mg} + \text{I}$



3.  $\text{Be} + \text{S}$



4.  $\text{Na} + \text{O}$



5.  $\text{Al} + \text{Br}$



## COVALENT BONDING

Name \_\_\_\_\_

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:



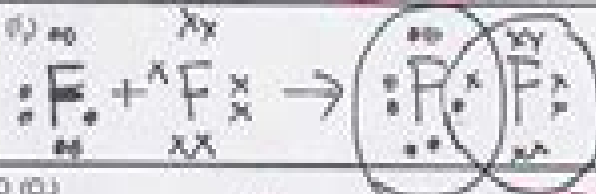
Note that hydrogen is content with 2, not 8, electrons.

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1.  $\text{H} + \text{H} (\text{H}_2)$



2.  $\text{F} + \text{F} (\text{F}_2)$



3.  $\text{O} + \text{O} (\text{O}_2)$



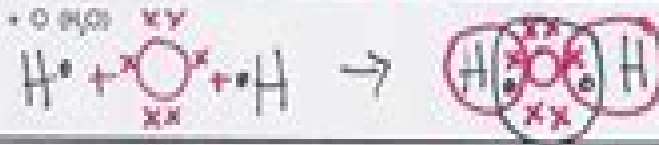
4.  $\text{N} + \text{N} (\text{N}_2)$



5.  $\text{C} + \text{O} (\text{CO})$



6.  $\text{H} + \text{O} (\text{H}_2\text{O})$



# Chapter 8 Covalent Bonding Workbook Answers

**Tao Wei**



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In a digitally-driven world where monitors reign great and quick communication drowns out the subtleties of language, the profound secrets and mental nuances concealed within phrases usually go unheard. However, set within the pages of **Chapter 8 Covalent Bonding Workbook Answers** a charming literary prize pulsating with natural feelings, lies an exceptional journey waiting to be undertaken. Written by an experienced wordsmith, this marvelous opus encourages readers on an introspective journey, delicately unraveling the veiled truths and profound impact resonating within the fabric of every word. Within the emotional depths of this emotional review, we shall embark upon a genuine exploration of the book's core themes, dissect its captivating publishing fashion, and fail to the powerful resonance it evokes deep within the recesses of readers' hearts.

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