Chapter 9 Practice Test - Naming and Writing Chemical Formulas

Matching

Match each itme with the correct statement below.

		Match each item with the correct statement I	well-read		
		a. monatomic ion	f.	cation	
		b. acid	8-	binary compound	
		c. base	h.	anion	
		d. law of definite proportions e. law of multiple proportions	1.	polyatomic ion	
	40	consists of a single atom with a positive or n			
	1.				
_	2.	atom or group of atoms having a negative ch			
	3.	atom or group of atoms having a positive cha	urge		
_	4.	tightly-bound group of atoms that behaves as a unit and carries a net charge			
	5.	compound composed of two different element	nts		
	6.	produces a hydrogen ion when dissolved in v	water		
	7.	produces a hydroxide ion when dissolved in	water		
	8.	In any chemical compound, the masses of el-	ement	s are always in the same proportion by mass.	
	9.	when two elements form more than one compound, the masses of one element that combine with the same mass of the other element are in the ratio of small, whole numbers			
		Thoice			
ldenti	fy the	choice that best completes the statement or a	rior where:	s the question.	
	10.	When naming a transition metal ion that can have more than one common ionic charge, the numerical value of the charge is indicated by a			
		a. prefix	42.3	Roman numeral following the name	
		b. suffix	dL	superscript after the name	
_	H.	Which of the following correctly provides the name of the element, the symbol for the ion, and the name of the ion?			
		a. fluorine, F*, fluoride ion	C	copper, Cu*, cuprous ion	
		 b. zinc, Zn^{2*}, zincate ion 	-dL	sulfur, S2-, sulfurous ion	
	12.	The nonmetals in Groups 6A and 7A a. lose electrons when they form ions b. have a numerical charge that is found by subtracting 8 from the group number			
		 nave a numerical charge that is found by all have ions with a "I charge 	SHOUR	acting 8 from the group number	
		d. end in -ate			
	13.	Which of the following is NOT a cation?			
		a. iron(III) ion	C.	Ca ²⁺	
		b. sulfate	d.	mercurous ion	

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