

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

| | |
|---|-------------------------|
| _____ lattice energy _____ | a. ionic bond |
| _____ bond energy _____ | b. covalent bond |
| _____ enthalpy of vaporization _____ | c. metallic bond |

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

| | | | |
|-----------------------------------|-----------------|--------------------------------------|---------------|
| _____ ionic _____ | a. H—F | _____ nonpolar covalent _____ | d. H—H |
| _____ ionic _____ | b. Na—Cl | _____ polar covalent _____ | e. H—C |
| _____ polar covalent _____ | c. H—O | _____ polar covalent _____ | f. H—N |

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Chapter 8 Covalent Bonding Answers Key

Arshad Iqbal



Chapter 8 Covalent Bonding Answers Key:

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