

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4.  $\text{H}_2\text{S}$  and  $\text{H}_2\text{O}$  have similar structures and their central atoms belong to the same group. Yet  $\text{H}_2\text{S}$  is a gas at room temperature and  $\text{H}_2\text{O}$  is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in  $\text{H}_2\text{O}$  bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Chapter 8 Covalent Bonding Worksheet Answer Key

**Shirley Soto**



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*Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know* Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more

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web cinderella by k i d s i n c o based on the story cinderella by brothers grimm the classic tale of cinderella who aided by a fairy godmother breaks away from the cruelty of her step sisters sisters to go to the palace ball where she enchants the prince all looks lost when she must return home at the stroke of midnight before the prince learns

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web cinderella readers theater play script authors elaine lindy country germany genre fairytales characters cinderella stepsister 1 stepsister 2 stepmother king queen prince short bow young lady ballroom music starts up prince and stepsister 1 start to dance to the center of the stage after a few

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web cinderella okay fairy godmother thank you scene 1 at cinderella s home narrator once upon a time there was a beautiful young girl named cinderella she had two ugly stepsisters who were very cruel to her stepsister 1 did you clean the kitchen cinderella yes i did stepsister 2 did you polish my shoes cinderella yes i did

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web cordelia cinderella s stepsister can be effectively played by a boy or a girl ursula the other stepsister can also be effectively played by a boy or a girl stepmother cinderella s stepmother the mice hector carries around a bag

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web prince many little birds narrator one once upon a time there was a beautiful woman who was married to a rich man one winter she fell gravely ill and she called her only daughter to her bedside cinderella s mother always be a good girl and i will look down from heaven and watch over you cinderella i promise i will mother