#### CHAPTER 7 REVIEW

## Chemical Formulas and Chemical Compounds

#### SECTION 7-1

#### SHORT ANSWER Answer the following questions in the space provided.

- In a Stock name such as iron(III) sulfate, the roman numeral tells us \_\_\_\_\_
  - (a) how many atoms of Fe are in one formula unit
  - (b) how many sulfate ions can be attached to the iron atom.
  - (c) the charge on each Fe ion
  - (d) the total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to \_\_\_\_\_\_
  - (a) change the number of moles represented by the formula
  - (b) change the charges on the other ions in the compound
  - (c) change the formula so that it no longer represents the compound it previously represented
  - (d) have no effect on the formula
- The explosive TNT has the molecular formula C<sub>2</sub>H<sub>3</sub>(NO<sub>2</sub>)<sub>3</sub>.
  - 4 elements a. How many elements make up this compound?
  - 6 oxygen atoms b. How many oxygen atoms are present in one molecule of  $C_2H_3(NO_2)_3$ ?
  - 21 atoms c. How many atoms in total are present in one molecule of C<sub>7</sub>H<sub>5</sub>(NO<sub>2</sub>)<sub>3</sub>?
  - 4.2 × 10<sup>24</sup> atoms d. How many atoms are present in a sample of 2 × 10<sup>23</sup> molecules of C<sub>7</sub>H<sub>2</sub>(NO<sub>2</sub>)<sub>3</sub>?
- 4. How many atoms are present in each of these formula units?
  - 11 atoms a, Ca(HCO<sub>3</sub>)<sub>2</sub>
    - 45 atoms b. C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>
  - 10 atoms c. Fe(ClO<sub>2</sub>)<sub>3</sub>
  - 9 atoms d. Fe(ClO<sub>3</sub>)<sub>2</sub>
- 5. N<sub>2</sub>Q<sub>5</sub> a. What is the formula for the compound dinitrogen pentoxide?
  - carbon(IV) sulfide b. What is the Stock name for the covalent compound CS<sub>2</sub>?
    - H<sub>2</sub>SO<sub>3</sub> c. What is the formula for sulfurous acid?
  - phosphoric acid d. What is the name for the acid H<sub>3</sub>PO<sub>4</sub>?

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# **Chapter 9 Chemical Names And Formulas Section Review Answers**

Morris Hein, Susan Arena

#### **Chapter 9 Chemical Names And Formulas Section Review Answers:**

Addison-Wesley Chemistry .1990 Essentials of Chemistry Dennis D. Staley, Antony C. Wilbraham, 1984 Prentice Introductory Chemistry Victor S. Krimsley, Darold E. Skerritt, Beverly B. Harrison, 1986 This Hall Chemistry .2000 pedagogically rich text has all the necessary features to hook introductory students and keep them interested and successful in preparatory chemistry This book carefully guides beginning students through the fundamental principals and calculations required for their subsequent success in either the general chemistry course or the short organic and biochemistry course Krimsley while dedicated to conceptual understanding and skill building presents a solid book that provides students with complete explanations on every point and helps them work through the material methodically with many examples and hints His approach is cohesive and coherent always patient and interesting Krimsley introduces all topics through an example students are already familiar with He continually reminds them of objectives and provides many opportunities to practice and then check their answers The text begins with a study of atomic and molecular structure prior to treating the various categories of chemical reactions The organization is designed to get students into chemistry quickly yet methodically The classification of matter begins in Chapter 2 before the chapter on measurement and the coverage of bonding appears in Chapter 8 The elementary concepts of chemistry are presented with an emphasis on mathematical calculations and problem solving strategies **Holt Chemistry** Ralph Thomas Myers, 2004 **Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity Fundamentals of Chemistry Ralph A. Burns, 2003 For one semester preparatory chemistry

courses or general purpose introductory chemistry courses This clearly written well illustrated versatile textbook provides thorough coverage of chemistry with a balance of problem solving skills real world applications and an emphasis on critical thinking and the process of science A supporting theme throughout the text continually emphasizes that chemistry is **Chemistry** James E. Brady, Fred Senese, 2004-02-04 Publisher Description **Foundations of College Chemistry** Morris Hein,1991 Features strong problem solving and carefully constructed problem sets Marketing Machinery, 1907 Machinery Lester Gray French, 1907 Basic Chemistry Steven S. Information Guide .1963 Zumdahl,1999 For a full description see catalog entry for Zumdahl Introductory Chemistry A Foundation 4 e Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis Railroad Gazette ,1894

#### Chapter 9 Chemical Names And Formulas Section Review Answers Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the energy of words has be much more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such may be the essence of the book **Chapter 9 Chemical Names And Formulas Section Review Answers**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall affect readers.

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Skill Practice 1 Classify the following as chemical changes (C) or physical changes (P) Given your answers to question 1
and the fact that this reaction takes place at 25oC Skill Practice 23 2004 by Jason Neil. All rights reserved. Skill Practice
23. Name: Date: Hour: Draw Lewis structures for each of the following. 1. NO3. 1 2. CH4. Skill Practice 26 Skill
Practice 26. Name: Date: Hour: 1. What does it mean to say that a bond is polar? One of the atoms Skill Practice 16 -
Atomic Size Skill Practice 16. Atomic Size. Practice. Name: KEY. Date: Hour: 1. What force of attraction does the second
energy level of a phosphorus atom "feel" from the Skill Practice 13 Obtain permission for classroom use at
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