

9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- ☒ c. transition metal
- d. nonmetal

5. The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the lesser charge, and the suffix *-ic* to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

Chapter 9 Chemical Names Formulas Quiz Answer Key

William H. Brown, Thomas Poon



Chapter 9 Chemical Names Formulas Quiz Answer Key:

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Chapter 1 Matter and change Chapter 2 measurement and calculations Chapter 3 Atoms The building blocks of matter Chapter 4 Arrangement of electrons in atoms Chapter 5 The periodic law Chapter 6 Chemical bonding Chapter 7 Chemical formulas and chemical compounds Chapter 8 Chemical equations and reactions Chapter 9 Stoichiometry Chapter 10 Physical characteristics of gases Chapter 11 Molecular composition of gases Chapter 12 Liquids and solids Chapter 13 Solutions Chapter 14 Ions in aqueous solution and colligative properties Chapter 15 Acids and bases Chapter 16 Acid base titrations Chapter 17 Reaction energy and reaction kinetics Chapter 18 Chemical equilibrium Chapter 19 Oxidation reduction reactions Chapter 20 Chemical thermodynamics Chapter 21 Carbon and hydrocarbons Chapter 22 Other organic compounds Chapter 23 Nuclear chemistry Chapter 24 Biological and Food chemistry Chapter 25 Environmental chemistry

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Table of Contents Chapter 9 Chemical Names Formulas Quiz Answer Key

1. Understanding the eBook Chapter 9 Chemical Names Formulas Quiz Answer Key
 - The Rise of Digital Reading Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Chemical Names Formulas Quiz Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Personalized Recommendations
 - Chapter 9 Chemical Names Formulas Quiz Answer Key User Reviews and Ratings
 - Chapter 9 Chemical Names Formulas Quiz Answer Key and Bestseller Lists

5. Accessing Chapter 9 Chemical Names Formulas Quiz Answer Key Free and Paid eBooks
 - Chapter 9 Chemical Names Formulas Quiz Answer Key Public Domain eBooks
 - Chapter 9 Chemical Names Formulas Quiz Answer Key eBook Subscription Services
 - Chapter 9 Chemical Names Formulas Quiz Answer Key Budget-Friendly Options
6. Navigating Chapter 9 Chemical Names Formulas Quiz Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Chemical Names Formulas Quiz Answer Key Compatibility with Devices
 - Chapter 9 Chemical Names Formulas Quiz Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Highlighting and Note-Taking Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Interactive Elements Chapter 9 Chemical Names Formulas Quiz Answer Key
8. Staying Engaged with Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Chemical Names Formulas Quiz Answer Key
9. Balancing eBooks and Physical Books Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Chemical Names Formulas Quiz Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Setting Reading Goals Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Fact-Checking eBook Content of Chapter 9 Chemical Names Formulas Quiz Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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