

The combustion of a hydrocarbon compound according to the reaction that follows is given. The reaction is spontaneous thermodynamically, indicating the spontaneity of the reaction. The reaction is:

- a. The combustion is spontaneous and the reaction is exothermic, indicating that the reaction is thermodynamically spontaneous. It will happen at a sufficient rate to be considered a reaction. This reaction is not a chemical reaction, it is a physical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction.

The reaction is spontaneous and the reaction is exothermic, indicating that the reaction is thermodynamically spontaneous. It will happen at a sufficient rate to be considered a reaction. This reaction is not a chemical reaction, it is a physical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction.

$$\Delta G = \Delta H - T\Delta S = -10.7 \text{ kJ} - (-1.2 \text{ kJ/K}) = -9.5 \text{ kJ/K} < 0$$

Under these conditions, the reaction is spontaneous and the reaction is exothermic, indicating that the reaction is thermodynamically spontaneous. It will happen at a sufficient rate to be considered a reaction. This reaction is not a chemical reaction, it is a physical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction.

- b. The reaction is spontaneous and the reaction is exothermic, indicating that the reaction is thermodynamically spontaneous. It will happen at a sufficient rate to be considered a reaction. This reaction is not a chemical reaction, it is a physical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction.

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- c. The reaction is spontaneous and the reaction is exothermic, indicating that the reaction is thermodynamically spontaneous. It will happen at a sufficient rate to be considered a reaction. This reaction is not a chemical reaction, it is a physical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction.

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- d. The reaction is spontaneous and the reaction is exothermic, indicating that the reaction is thermodynamically spontaneous. It will happen at a sufficient rate to be considered a reaction. This reaction is not a chemical reaction, it is a physical reaction. It is a chemical reaction of the combustion of a hydrocarbon compound, indicating it is a chemical reaction.

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$$\Delta G = \Delta H + p\Delta V = \Delta H + p\Delta V = -10.7 + 1.2 = -9.5 \text{ kJ/K}$$

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**Yicheng Fang**



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