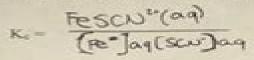
Report Sheet - Experiment 13-Chemical Equilibrium

- 1. Write the chemical equation for the reversible reaction studied in this experiment:

 Fe *(aq)+ SCN * > Fe (SCN) *(aq)
- 2. Write the equilibrium expression for the reaction studied in this experiment:



3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added Fe(NO ₁) ₁	Reddish color	dorker
2	Added KSCN	Reddish color	darker
3	Added NH ₄ CI	Reddish color	Lighter
4	Added AgNO	Reddish color	Lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	Fesch to a mix	Shifted Aykt	principle but react on from left to right accust was when more reactant add
3	[SCN] was	Shifted aight	as more reactant was added, the concentration of the product goes up
4	Fezz concentration was	Shifetell Just+	Fe COL WILL Form which leads to a lower Fe to concentration

Chem Fax Lab Answers Chemical Equilibrium

R Pring

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Fundamentals of Chemistry: Laboratory Studies Frank Brescia, 2012-12-02 Fundamentals of Chemistry Laboratory Studies Third Edition is a manual that provides instruction on techniques of chemical laboratory operations Each experiment is discussed in terms of the major objective the experimental approach to the objective the measurements or observations to be made and the calculation and interpretation of results Topics covered include manipulation weights and measures molecular weight acids and bases gravimetric and volumetric stoichiometry and thermochemistry This book is comprised of 43 chapters divided into 14 sections and begins by presenting general information on metric and other units common laboratory equipment and chemical laboratory methods. The first chapter introduces the reader to the Bunsen burner and the principles of glass working followed by a discussion on mass and volume measurements including the determination of density The following chapters focus on states of matter molecular weight stoichiometry and intermolecular forces Preparations and syntheses are also considered along with chemical equilibrium and electrochemistry The final section is devoted to qualitative analysis particularly of cations and anions This monograph is intended primarily for students of A Laboratory Manual of General Chemistry Leon Burr Richardson, Andrew Jackson Scarlett, 1929 chemistry The American Catalogue ,1905 A Laboratory outline of general chemistry Alexander Smith, 1899 Chemical The Journal of Industrial and Engineering Chemistry ,1914 Abstracts ,1910 Register Cornell University, 1892 "The" Cornell University register Cornell University, 1893 The American Physician ,1925 The Education Index Chemical News and Journal of Physical Science, 1895 .1986 **Chemical News and Journal of Industrial Science** Books in Print Supplement ,2002 Chemical Engineering Catalog ,1919 ,1895 A Dictionary of Applied **Chemistry** Sir Thomas Edward Thorpe,1921 A Laboratory Outline of College Chemistry Alexander Smith, 1916 **Current Index to Journals in Education**, 1985 Laboratory Manual of General Chemistry Arthur Becket Lamb, 1916 Journal of Chemical Education, 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers The Chemical Trade Journal and Chemical Engineer, 1908

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