

Section 7.5: Quantitative Changes in Equilibrium Systems

Tutorial 1 Practice, page 452

1. (a) Given: $[\text{HI}(\text{g})] = 0.14 \text{ mol/L}$; $[\text{H}_2(\text{g})] = 0.040 \text{ mol/L}$; $[\text{I}_2(\text{g})] = 0.010 \text{ mol/L}$;

$K = 0.020$

Required: Q ; direction of reaction to reach equilibrium

$$\begin{aligned}\text{Solution: } Q &= \frac{[\text{H}_2(\text{g})][\text{I}_2(\text{g})]}{[\text{HI}(\text{g})]^2} \\ &= \frac{(0.040)(0.010)}{(0.14)^2}\end{aligned}$$

$$Q = 0.020$$

Statement: The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium.

(b) Given: $[\text{HI}(\text{g})] = 0.20 \text{ mol/L}$; $[\text{H}_2(\text{g})] = 0.15 \text{ mol/L}$; $[\text{I}_2(\text{g})] = 0.090 \text{ mol/L}$; $K = 0.020$

Required: Q ; direction of reaction to reach equilibrium

$$\begin{aligned}\text{Solution: } Q &= \frac{[\text{H}_2(\text{g})][\text{I}_2(\text{g})]}{[\text{HI}(\text{g})]^2} \\ &= \frac{(0.15)(0.090)}{(0.20)^2}\end{aligned}$$

$$Q = 0.34$$

Statement: The value of Q is 0.34. Q is greater than K , so the product concentrations are greater than at equilibrium; the reaction will shift toward the left, more HI.

2. Time 1:

Given: $[\text{N}_2\text{O}_4(\text{g})] = 0.80 \text{ mol/L}$; $[\text{NO}_2(\text{g})] = 1.55 \text{ mol/L}$; $K = 0.87$

Required: Q ; direction of reaction to reach equilibrium

Solution: $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$

$$\begin{aligned}Q &= \frac{[\text{NO}_2(\text{g})]^2}{[\text{N}_2\text{O}_4(\text{g})]} \\ &= \frac{(1.55)^2}{0.80}\end{aligned}$$

$$Q = 3.0$$

Statement: The value of Q is 3.0. Q is greater than K , so the product concentrations are greater than at equilibrium; the reaction will shift toward the left, more N_2O_4 .

Time 2:

Given: $[\text{N}_2\text{O}_4(\text{g})] = 1.66 \text{ mol/L}$; $[\text{NO}_2(\text{g})] = 1.20 \text{ mol/L}$; $K = 0.87$

Required: Q ; direction of reaction to reach equilibrium

$$\begin{aligned}\text{Solution: } Q &= \frac{[\text{NO}_2(\text{g})]^2}{[\text{N}_2\text{O}_4(\text{g})]} \\ &= \frac{(1.20)^2}{1.66}\end{aligned}$$

$$Q = 0.867$$

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Darrell D. Ebbing



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