

Name Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 , is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$\frac{99.2 \text{ mol } H_2}{2 \text{ mol } H_2} \times \frac{-484 \text{ kJ}}{1} = -24,006 \text{ kJ}$$

$-24,006 \text{ kJ}$
 ~~$-24,006 \text{ kJ}$~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?

reverse so
change sign of
 ΔH out in 2
quantity in 1/2



Exothermic

3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

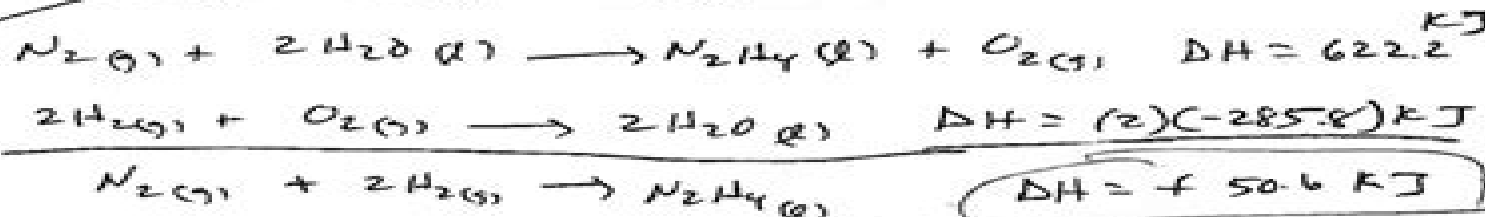
$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_8O} \times 4.5 \text{ mol } C_3H_8O = -1920.6 \text{ kcal}$$

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



Reverse
double
quantity

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1. Understanding the eBook Chemistry Practice Problems And Answers
 - The Rise of Digital Reading Chemistry Practice Problems And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Problems And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practice Problems And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Problems And Answers

- Personalized Recommendations
 - Chemistry Practice Problems And Answers User Reviews and Ratings
 - Chemistry Practice Problems And Answers and Bestseller Lists
5. Accessing Chemistry Practice Problems And Answers Free and Paid eBooks
 - Chemistry Practice Problems And Answers Public Domain eBooks
 - Chemistry Practice Problems And Answers eBook Subscription Services
 - Chemistry Practice Problems And Answers Budget-Friendly Options
 6. Navigating Chemistry Practice Problems And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Practice Problems And Answers Compatibility with Devices
 - Chemistry Practice Problems And Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practice Problems And Answers
 - Highlighting and Note-Taking Chemistry Practice Problems And Answers
 - Interactive Elements Chemistry Practice Problems And Answers
 8. Staying Engaged with Chemistry Practice Problems And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practice Problems And Answers
 9. Balancing eBooks and Physical Books Chemistry Practice Problems And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practice Problems And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Practice Problems And Answers
 - Setting Reading Goals Chemistry Practice Problems And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Practice Problems And Answers

- Fact-Checking eBook Content of Chemistry Practice Problems And Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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