

# Unit 1 – Atomic Structure MC

## Regents Questions Key

3

1. Which element has an atom with the electron configuration 2-8-8-2?

- |       |       |
|-------|-------|
| 1. Mg | 3. Ca |
| 2. Ni | 4. Ge |

3 An atom of the element has 20 electrons ( $2 + 8 + 8 + 2$ ) and 20 protons. Use the Periodic Table of the Elements and find the electron configuration of the element whose atomic number is 20.

1

2. Which type of radiation has *neither* mass nor charge?

- |            |          |
|------------|----------|
| 1. gamma   | 3. alpha |
| 2. neutron | 4. beta  |

1 All forms of electromagnetic radiation, such as gamma radiation, have neither mass nor electric charge.

3

3. Which list of particles is in order of increasing mass?

- (1) proton → electron → alpha particle
- (2) proton → alpha particle → electron
- (3) electron → proton → alpha particle
- (4) alpha particle → electron → proton

3 An electron is the least massive of the three particles; a proton is approximately 2,000 times more massive than an electron; an alpha particle is the nucleus of a helium-4 atom and is approximately 4 times more massive than a proton.

3

4. Compared to a sodium atom in the ground state, a sodium atom in the excited state must have

- |                                  |                                    |
|----------------------------------|------------------------------------|
| 1. a greater number of electrons | 3. an electron with greater energy |
| 2. a smaller number of electrons | 4. an electron with less energy    |

3 Atoms become excited by absorbing energy. Therefore, an excited atom has more energy than it had in the ground state.

**Wrong Choices Explained:**

(1), (2) An excited atom is still a neutral atom. It has neither greater nor fewer electrons than an atom in the ground state.

1

5. Which particles are isotopes of each other?

# Chemistry Regents Questions And Answers Atomic Structure

**Barron's Educational Series, Albert S.  
Tarendash**



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