

12.2

CHEMICAL CALCULATIONS

Section Review

Objectives

- Construct mole ratios from balanced chemical equations and apply these ratios in mole-mole stoichiometric calculations
- Calculate stoichiometric quantities from balanced chemical equations, using units of moles, mass, representative particles, and volumes of gases at STP

Key Equations

- mole-mole relationship used in every stoichiometric calculation:



(given quantity) (wanted quantity)

- $x \text{ mol } G \times \frac{b \text{ mol } W}{a \text{ mol } G} = \frac{xb}{a} \text{ mol } W$

Given

Mole Ratio

Calculated

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

- Mole ratios from balanced equations may be used to solve _____ 1. _____
- problems with other units such as numbers of _____ 1 _____ and _____ 2 _____ 2. _____
- of gases at STP. The _____ 3 _____ from the balanced equation are used _____ 3. _____
- to write conversion factors called _____ 4 _____. These conversion factors _____ 4. _____
- are used to calculate the numbers of moles of _____ 5 _____ from a given _____ 5. _____
- number of moles of _____ 6 _____. In mass-mass calculations, the molar _____ 6. _____
- mass is used to convert mass to _____ 7 _____. _____ 7. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 8. In mass-mass calculations, the molar mass is used to convert mass to moles.
- _____ 9. The mole ratio 2 mol HF/1 mol SnF₂ can be used to determine the mass of SnF₂ produced according to the equation:
- $$\text{Sn}(s) + 2\text{HF}(g) \rightarrow \text{SnF}_2(s) + \text{H}_2(g)$$

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1

Albert S. Tarendash



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