

Section 11.4 Empirical and Molecular Formulas

In your textbook, read about percent composition.

Answer the following questions.

1. What is the percent composition of a compound?

The percent of the total mass of the compound that is due to that component.

2. Describe how to find the percent composition of a compound if you know the mass of a sample of a compound and the mass of each element in the sample.

÷ the mass of each element in the sample by the mass of the sample. × each by 100.

In your textbook, read about empirical and molecular formulas.

Circle the letter of the choice that best answers the question.

3. Which information about a compound can you use to begin to determine the empirical and molecular formulas of the compound?

a. mass of the compound

b. number of elements in the compound

c. percent composition of the compound

d. volume of the compound

4. You have determined that a compound is composed of 0.300 moles of carbon and 0.600 moles of oxygen. What must you do to determine the mole ratio of the elements in the empirical formula of the compound?

a. Multiply each mole value by 0.300 mol.

b. Multiply each mole value by 0.600 mol.

c. Divide each mole value by 0.300 mol.

d. Divide each mole value by 0.600 mol.

5. The mole ratio of carbon to hydrogen to oxygen in a compound is 1 mol C : 2 mol H : 1 mol O. What is the empirical formula of the compound?

a. CHO

b. CH₂O

c. C₂H₂O₂

d. C₂H₂O₃

6. You calculate the mole ratio of oxygen to aluminum in a compound to be 1.5 mol O : 1 mol Al. What should you do to determine the mole ratio in the empirical formula of the compound?

a. Multiply each mole value by 1.5.

b. Multiply each mole value by 2.

c. Divide each mole value by 1.5.

d. Divide each mole value by 2.

7. What is the relationship between the molecular formula and the empirical formula of a compound?

a. (molecular formula)/(empirical formula) = n

b. molecular formula = $\frac{\text{empirical formula}}{n}$

c. molecular formula = (empirical formula) n

d. molecular formula = $\frac{n}{\text{empirical formula}}$

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Zumdahl



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Resources in Education ,1980 Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl,Paul B. Kelter,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples **Glencoe Chemistry** Glencoe/McGraw-Hill,2002

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Principles of Modern Chemistry Wade A. Freeman,David W. Oxtoby,1996 PRINCIPLES OF MODERN CHEMISTRY has dominated the honors and high mainstream general chemistry courses and is considered the standard for the course The fifth

edition is a substantial revision that maintains the rigor of previous editions but reflects the exciting modern developments taking place in chemistry today Authors David W Oxtoby and H P Gillis provide a unique approach to learning chemical principles that emphasizes the total scientific process from observation to application placing general chemistry into a complete perspective for serious minded science and engineering students Chemical principles are illustrated by the use of modern materials comparable to equipment found in the scientific industry Students are therefore exposed to chemistry and its applications beyond the classroom This text is perfect for those instructors who are looking for a more advanced general chemistry textbook

Chemistry Study Guide and Student Solutions Guide Sixth Edition Zumdahl,2003-01-01 **Student Study Guide and Solutions Manual to accompany Organic Chemistry** David R. Klein,2010-06-28 This text is an unbound binder ready edition Students often say I studied 40 hours for this exam and I still didnt do well Where did I go wrong Most instructors hear this complaint every year In many cases it is true that the student invested countless hours only to produce abysmal results Often inefficient study habits are to blame The important question is why do so many students have difficulty preparing themselves for organic chemistry exams There are certainly several factors at play here but perhaps the most dominant factor is a fundamental disconnect between what students learn and the tasks expected of them To address the disconnect in organic chemistry instruction David Klein has developed a textbook that utilizes a skills based approach to instruction The textbook includes all of the concepts typically covered in an organic chemistry textbook but special emphasis is placed on skills development to support these concepts This emphasis upon skills development will provide students with a greater opportunity to develop proficiency in the key skills necessary to succeed in organic chemistry As an example resonance structures are used repeatedly throughout the course and students must become masters of resonance structures early in the course Therefore a significant portion of chapter 1 is devoted to drawing resonance structures Two chapters 6 and 12 are devoted almost entirely to skill development Chapter 6 emphasizes skills that are necessary for drawing mechanisms while chapter 12 prepares the student for proposing syntheses In addition each chapter contains numerous Skillbuilders each of which is designed to foster a specific skill Each skillbuildercontains three parts 1 Learn the Skill a solved problem that demonstrates a particular skill 2 Practice the Skill numerous problems similar to the solved problem that give the students an opportunity to practice and master the skill 3 Apply the Skill one or two more challenging problems in which the student must apply the skill in a slightly different environment These problems include conceptual cumulative and applied problems that encourage students to think out of the box Sometimes problems that foreshadow concepts introduced in later chapters are also included AllSkillBuilders are visually summarized at the end of each chapter Skillbuilder review followed by a list of suggested in chapter and end of chapter practice problems David Klein is a lecturer at Johns Hopkins University where he teaches Organic and General Chemistry He is a dynamic and creative teacher and uses analogy to help students grasp difficult topics Kleins unique informal voice and manner of presentation help

students truly master key topics in this course He is also the author of Organic Chemistry as a Second Language and General Chemistry as a Second Language which have both been highly successful **Chemistry Lesson Plans, Study Guides, and Lecture Notes** Doctor Chemistry, 2017-08-13 The Ultimate Guide to Learning or Teaching Chemistry This book contains the real lecture notes and slide of a highly effective high school and college Chemistry teacher Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Atomic Theory Periodic Table VSEPR Ionic and Covalent Bonding Geometric Bonding The Mole and Molar Mass Equation Balancing Thermodynamics Stoichiometry States of Matter Gas Laws and Calculations Reaction Calculations Acids and Bases Limiting Reagents Redox and Electro Chemistry Organic Chemistry Basics **Concepts of General, Organic, and Biological Chemistry**, 1981

Student Study Guide/Solutions Manual to accompany General, Organic & Biological Chemistry Janice Smith, Erin Smith Berk, 2009-02-19 Each chapter of the Student Study Guide begins with a chapter review tied to the chapter goals in the text Next Sample problems are supplied and stepped out through the solution for each type of problem covered in the chapter A Self Test serves up fill in the blank exercises to assess learning with answers supplied at the end of the chapter Finally chapters end with the solutions for all of the in chapter problems as well as for the odd numbered end of chapter problems

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