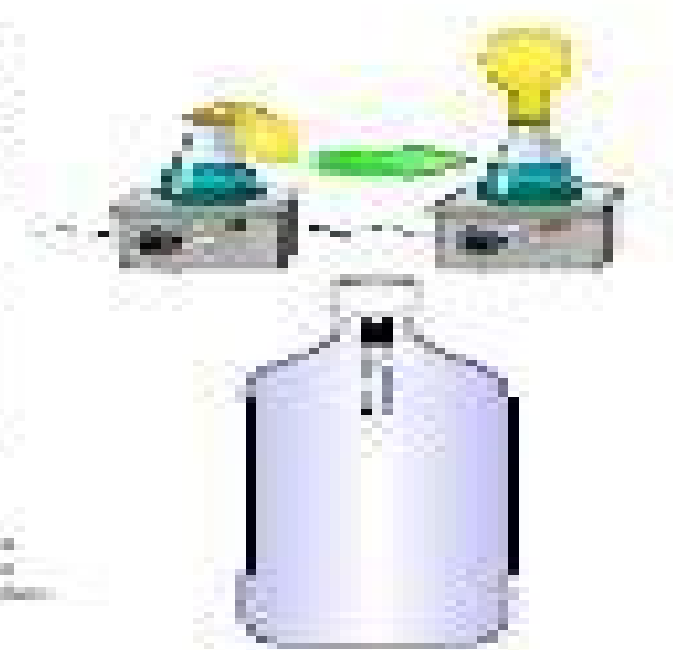


Using the Combined Gas Law

We believe in gaseous molecules only fixed and the atmosphere is a free place, the balloon floats up as it heats up. This happens even though we still have the same gas in the balloon. As the temperature increases, the gas expands, pushing back the air. This is what causes the air to rise. The air rises because the pressure is greater than the air pressure. The air pressure is lower than the air pressure in the balloon. The air pressure is lower than the air pressure in the balloon. The air pressure is lower than the air pressure in the balloon.



The Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Where:

P_1 = initial pressure

V_1 = initial volume

T_1 = initial temperature (K)

P_2 = final pressure

V_2 = final volume

T_2 = final temperature (K)

This equation can be used to find the final pressure, volume, or temperature of a gas if you know the initial pressure, volume, or temperature. If you know the initial pressure, volume, or temperature, you can find the final pressure, volume, or temperature.

Example Problem

A gas with a volume of 100 mL at 100 kPa and 300 K is heated to 400 K and the pressure is increased to 150 kPa. What is the final volume?

$$P_1 = 100 \text{ kPa}, T_1 = 300 \text{ K}, P_2 = 150 \text{ kPa}, T_2 = 400 \text{ K}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$V_2 = \frac{P_1 V_1 T_2}{P_2 T_1} = \frac{(100 \text{ kPa})(100 \text{ mL})(400 \text{ K})}{(150 \text{ kPa})(300 \text{ K})} = 88.9 \text{ mL}$$

Chemistry Form Ws6 4 1a Answers

Wolfgang Guggemos

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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