

SEMESTER 1 FINAL REVIEW – CHEMISTRY

2015-16

Please ensure that you know how to answer these questions correctly, these are all good example problems for your final exam. Study your notes, labs, homework assignments and use the PowerPoint presentations that are online. Any of this information is fair game to be asked on the final exam. Good luck and ask questions in class during the review times if you have them. Don't wait to the last minute to start reviewing!

1. What is the number of significant figures in 1007000.0
2. What is the difference between accuracy and precision?
3. What is Avogadro's number? Why is it useful?
4. Find the molecular weight of CO_2 , NH_3 , SF_6 ?
5. Be able to use density to calculate mass and or volume for various objects. Also, understand that density can be used at times with moles and particles through the $D=m/v$ formula.
6. What is a mole? Be able to explain how the mole is a central connecting idea between mass, particles and gases @STP
7. What is the molar volume of a gas at STP?
8. What are the seven diatomic elements?
9. Define: atom, molecule, diatomic gas, monoatomic gas, element, compound, STP
10. What elements form the following bonds: ionic, covalent
11. Give an example of the following (both name and formula)
 - a. an ionic compound
 - b. a covalent compound
 - c. a metal
 - d. a nonmetal
 - e. a diatomic gas
12. Name the following:

a) NaOH	b) SO_2	c) C_2SO_4	d) HCl	e) $\text{Ca}(\text{NO}_3)_2$	f) NH_4
g) N_2O_5	h) $\text{Al}_2(\text{SO}_4)_3$	i) K_3PO_4	j) Na_2O	k) CO	l) CO_2
13. Write formulas for following
 - a) dinitrogen trioxide
 - b) potassium sulfate
 - c) hydrogen phosphate
 - d) sodium bromide
 - e) ammonium acetate
 - f) manganese IV nitrate
 - g) iron III hydroxide
 - h) aluminum carbonate
14. Write the following in correct scientific notation:

a) 476	b) 0.0000000002223	c) 3240000000000000
d) 462.9×10^{23}	e) $0.000000345 \times 10^{12}$	
f) 5677.9×10^{-12}	g) $0.000000223 \times 10^{-6}$	
15. How many oxygen atoms are there in 420 grams of carbon dioxide?
16. What is the mass of 1.2×10^{24} molecules of bromine?
17. What is the volume occupied by 243 grams of oxygen gas at STP?
18. How do you convert from:

a) moles \rightarrow grams	b) grams \rightarrow moles
c) moles \rightarrow molecules	d) moles \rightarrow liters of gas at STP
e) moles \rightarrow liters of solution	

Chemistry Semester 1 Exam Answer Key

Mary Millhollon, Richard H. Langley



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Text Book of Physical Chemistry Semester-IV, Course-10 by BVR VENKATA RAO BASA (BVR SIR),2025-03-12 Text Book of Physical Chemistry Semester IV Course 10 by BVR is meant for BSc students pursuing Organic Chemistry Honors The syllabus may vary slightly from college to college The topics covered are dilute solutions electrochemistry phase rule photochemistry Practicals CST and conductometry Objective questions for entrance exams included and also video links for further knowledge included Integrating Sustainable Development into the Curriculum Enakshi Sengupta,Patrick Blessinger,Taisir Subhi Yamin,2020-03-18 This book explores the value of institutions of higher education in leading the way on the topic of sustainability education by ensuring that it is well entrenched in the curriculum as well as everyday practice and lifestyles Nursing School Entrance Exam Review William H. Brown,Frederick A. Bettelheim,Jerry March,2004-04

Artificial Intelligence in Education. Posters and Late Breaking Results, Workshops and Tutorials, Industry and Innovation Tracks, Practitioners, Doctoral Consortium and Blue Sky Andrew M. Olney,Irene-Angelica Chounta,Zitao Liu,Olga C. Santos,Ig Ibert Bittencourt,2024-07-01 This volume constitutes poster papers and late breaking results presented during the 25th International Conference on Artificial Intelligence in Education AIED 2024 which took place in Recife Brazil during July 8 12 2024 The 18 full papers and 92 short papers were carefully reviewed and selected from 200 submissions They are organized in topical sections as follows Part One Blue Sky Industry Innovation and Practitioner WideAIED and Late Breaking Results Part Two Late Breaking Results Doctoral Consortium Workshops and Tutorials **The Publishers Weekly** ,1967 **El-Hi Textbooks in Print** ,1977 **Student-Assisted Teaching** Judith E. Miller,James E. Groccia,Marilyn Sue Miller,Marilyn S. Miller,2001-01-15 This book provides a range of models for undergraduate student assisted teaching partnerships to help teachers and administrators make learning more student centered effective and productive The 31 models describes a range of approaches and applications in a variety of settings and disciplines The chapters are 1 Establishing a Common Ground a Conjoint Training Model for Instructors and Peer Educators Eve M Adams Susan C Brown and Terry L Cook 2 Lessons from Peers The Design Exchange Mark J Chidister Frank H Bell Jr And Kurt M Earnest 3 Peer Teaching in the Experimental College Robyn Gittleman and Howard Woolf 4 Peer Facilitators as Lead Freshman Seminar Instructors Jean M Henscheid 5 The Teaching Teams Program a Just in Time Model for Peer Assistance Harold P Larson Reed Mencke Stacy J Tollefson Elizabeth Harrison and Elena Merman 6 The Teaching Teams Program Transforming the Role of the Graduate Teaching Assistant David A Wood Jr Jennifer L Hart Stacy J Tollefson Dawn E DeToro and Julie Libarkin 7 The Teaching Teams Program Empowering Undergraduates in a Student Centered Research University Lacey A Stover Kirstin A Story Amanda M Skousen Cynthia E Jacks Heather Logan and Benjamin T Bush 8 Peer Assisted Cooperative Learning An Experiment in Educational Quality and Productivity Judith E Miller David DiBiasio John Minasian and James S Catterall 9 Students Managing to Learn Teachers Learning To Manage Martin H Murray 10

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 Contains 18 figures 59 tables and 178 references Sld **Reports from the Classroom** Sarah H. Huyvaert, 1995

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Test Bank Jackie Fellows, James Henslin, 2003

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