- A piece of aluminum foil 1.00 cm square and 0.550 mm thick is allowed to react with bromine to form aluminum bromide.
 - a.) How many moles of aluminum were used? (density Al = 2.699 g/cm²)

b.) How many grams of aluminum bromide form, assuming the aluminum reacts completely? What type of reaction is this? ZAI +3Br, - ZAIBG

5.50×10 mol A1 ×
$$\frac{2 \text{ mol AlBe}_3}{2 \text{ mol He}}$$
 × $\frac{2 \text{ lobe lefty AlBe}_3}{1 \text{ mol HeBe}_3}$ = $\frac{1.47+9}{2 \text{ composition reaction}}$

a. Determine the percent composition of H and C for benzene

b. Write a balanced combustion reaction for benzene.

c. Assuming you have 364g of benzene, how many moles of O2 will be required to completely combust this amount of benzene?

d. How many grams of CO₂ will this produce? How many grams of H₂O?

Cb₇:
$$364gC_0H_0 \times \frac{1mol GH_0}{78.12g}C_0H_0 \times \frac{12mol Co_2}{2mol GH_0} \times \frac{44.04gCo_2}{1 mol Co_2} = 1.23 \times 10^3 g Co_2$$

H₂0: $344gC_0H_0 \times \frac{1mol C_0H_0}{78.12gC_0H_0} \times \frac{6mol H_2O}{2mol GH_0} \times \frac{16 H_2O}{1 mol H_2O} = 3.53 \times 10^3 g H_2O$

e. You perform the combustion, and 956g of CO₂ are produced. What is the percent yield?

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Xiang Xie

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