

8. How many moles of copper atoms are needed to supply the same number of atoms as does 235 grams of iron? Key.
- $$235 \text{ g} \left( \frac{1 \text{ mol}}{55.8 \text{ g}} \right) = 4.21 \text{ moles of Fe atoms}$$
- 4.21 moles of Cu atoms would have the same number of atoms.
9. How many grams of carbon are needed to supply twice as many atoms as does 27 grams of beryllium?
- $$27 \text{ g} \left( \frac{1 \text{ mol}}{9.0 \text{ g}} \right) = 3.0 \text{ moles Be atoms} \quad \text{twice atoms } 2(3.0 \text{ moles}) = 6.0 \text{ moles}$$
- $$6.0 \text{ moles C} \left( \frac{12.0 \text{ g}}{1 \text{ mol}} \right) = 72 \text{ g of carbon.}$$
10. 62.0 grams of rust is composed of 43.4 grams of iron. Determine the empirical formula for rust.
- $$43.4 \text{ g Fe} \rightarrow 43.4 \text{ g} \div 55.85 \text{ g/mol} = 0.78 \text{ mol Fe} \quad \text{mol ratio: Fe : O}$$
- $$62.0 - 43.4 \text{ g O} \rightarrow 18.6 \text{ g} \div 16.00 \text{ g/mol} = 1.16 \text{ mol O} \quad 2 : 3$$
- $\text{Fe}_2\text{O}_3$
11. A solution of NaCl contains 35 g of solute dissolved in 100g of water. What is the molarity of the solution?
- $$\text{FW NaCl} = 58.5 \text{ g/mol} \quad 100 \text{ g H}_2\text{O} \rightarrow 100 \text{ mL H}_2\text{O} \rightarrow 0.100 \text{ L}$$
- $$35 \text{ g} \div 58.5 \text{ g/mol} = 0.60 \text{ mol}$$
- $$M = \frac{n}{V} = \frac{0.60 \text{ mol}}{0.100 \text{ L}} = 6.0 \text{ M}$$
12. The following three containers have identical capacities. Each contains Avogadro's number of molecules of a gaseous element. Fill in the blanks with the missing information.



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