

Experiment 9

Acid-Base Titration

OBJECTIVES:

- To learn basic chemical characteristics of acids and bases.
- To learn how to perform titrations with precision.
- To understand how titrations can be used in chemical analysis.
- To prepare and standardize an NaOH solution.
- To use your standardized NaOH to determine the molarity of acetic acid in vinegar.

INTRODUCTION:

All Freshman Chemistry labs include at least one **titration** experiment. Titration is an old and time-honored method of **quantitative analysis**.¹ In every titration, we have some reaction occurring between some pair of reactants, A & B.

General



Reaction

A titration reaction usually falls into one of the three following categories:

(1) A Precipitation Reaction.



(2) An Acid-Base Reaction.

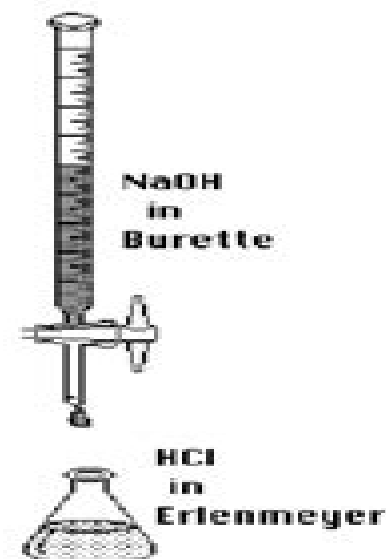


(3) An Oxidation-Reduction Reaction.



In today's experiment the reaction will be of type (2)—an acid-base neutralization reaction. As always in a titration, we place one reactant (a solution of NaOH) in a Burette and the other

¹ Quantitative Analysis techniques are designed to answer quantitative questions concerning how much of something is present in a sample. Quantitative Analysis techniques are to be distinguished from **Qualitative Analysis** techniques that are designed to answer qualitative questions concerning what substances are present in a sample.



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James Sherwood Fritz



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