

If it takes 50 mL of 0.5 M KOH solution to completely neutralize 125 mL of sulfuric acid solution (H_2SO_4), what is the concentration of the H_2SO_4 solution?



1 mole of KOH reacts with $\frac{1}{2}$ mole of H_2SO_4

$$n(\text{KOH}) = c \times V = 0.5 \times 0.05 = 0.025 \text{ mol} \leftarrow \text{KNOWN}$$

ACID-BASE

$$n(\text{H}_2\text{SO}_4) = \frac{1}{2} n(\text{KOH}) = 0.0125 \text{ mol} \leftarrow \text{UNKNOWN}$$

TITRATION

$$c(\text{H}_2\text{SO}_4) = \frac{n}{V} = \underline{\underline{0.1 \text{ mol L}^{-1} \text{ of } \text{H}_2\text{SO}_4}}$$

EXAMPLE

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 - The Rise of Digital Reading Acid Base Titration Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Titration Practice Problems With Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Acid Base Titration Practice Problems With Answers
 - Personalized Recommendations
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Titration Practice Problems With Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Titration Practice Problems With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Acid Base Titration Practice Problems With Answers
 - Carving Out Dedicated Reading Time

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 - Fact-Checking eBook Content of Acid Base Titration Practice Problems With Answers
 - Distinguishing Credible Sources
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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