



Name:

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Making Salts Worksheet

This worksheet accompanies *Making Salts Summary* and *Precipitation Reactions Summary*.

1. When an acid reacts with a base, a neutralisation reaction occurs, producing a salt. Using the following equation, name the salts that are produced in the acid-base reactions.

acid + base \longrightarrow salt

- a) hydrochloric acid + ammonia \longrightarrow
- b) nitric acid + sodium hydroxide \longrightarrow
- c) phosphoric acid + potassium hydroxide \longrightarrow



2. There are three other ways of making salts from acids. Complete the following equations to illustrate them.

- a) acid + metal \longrightarrow salt +
- b) acid + \longrightarrow salt + water
- c) acid + carbonate \longrightarrow salt + + water

3. Use the above information to complete the following reactions that form soluble salts using acids.

- a) iron oxide + sulfuric acid \longrightarrow + water
- b) + hydrochloric acid \longrightarrow magnesium chloride + H_2O + CO_2
- c) zinc + \longrightarrow zinc nitrate + hydrogen

4. Name two ways in which salt is used by humans in everyday life.

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Acid Base Or Salt Physical Science If8767 Answers

Brian J. Knapp



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Holt Science Spectrum Physical Science Chapter 9 Resource File: Acids, Bases, and Salts Holt Rinehart & Winston,2008-01-01 *Acids, Bases, and Salts* Brian J. Knapp,1998 *Acids and Bases* Kristi Lew,2009 Learn about acids and bases chemical components of the natural world that play key roles in medicine and industry *Acids, Bases, and Salts* Adams, This note is part of Quality testing Acids and Bases Brian G. Cox,2013-01-31 Acids and bases are ubiquitous in chemistry Our understanding of them however is dominated by their behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described

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