

ACID-BASE TITRATION

Topic: Acid-Base Titration

To determine the concentration of an acid (or base), we combine it with a known (or known) of known concentration and a colorfully neutralized. The point at which neutralization happens is the endpoint, is noted by the change in color of the solution. Also see the following equation:

$$M_1 V_1 = M_2 V_2 = M_3 V_3 \quad \text{where } M \text{ is molarity}$$

Solve the problems below:

1. A 25.0 mL sample of 0.100 M acetic acid is titrated with 15.0 mL of 0.100 M NaOH. What is the pH of the solution?
- 0.100 M

2. A 25.0 mL sample of 0.100 M acetic acid is titrated with 15.0 mL of 0.100 M NaOH. What is the pH of the solution?
- 0.100 M

3. A 25.0 mL sample of 0.100 M acetic acid is titrated with 15.0 mL of 0.100 M NaOH. What is the pH of the solution?
- 0.100 M

4. A 25.0 mL sample of 0.100 M acetic acid is titrated with 15.0 mL of 0.100 M NaOH. What is the pH of the solution?
- 0.100 M

5. A 25.0 mL sample of 0.100 M acetic acid is titrated with 15.0 mL of 0.100 M NaOH. What is the pH of the solution?
- 0.100 M

Acid Base Titration Instructional Fair Answers

Yijin Wang



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