



Electrons and Reactivity

Atoms contain

- a very small nucleus packed with neutrons and positively charged protons.
- a large volume of space around the nucleus that contains the negatively charged electrons.

It is the electrons that determine the physical and chemical properties of atoms.

Elements react in order to achieve the same electronic structure of the nearest noble gas

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Ionic and Covalent Bonds

Atoms form octets

- to become more stable.
- by losing, gaining, or sharing valence electrons.
- by forming ionic or covalent bonds.



Covalent Compounds

Covalent compounds are formed when nonmetals share valence electrons forming a covalent bond.

Covalent compounds consist of molecules, which are discrete groups of atoms, such as:

- water, H_2O 2 H atoms + 1 O atom
- carbon dioxide, CO_2 1 C atom + 2 O atoms
- glucose, $\text{C}_6\text{H}_{12}\text{O}_6$ 6 C atoms + 12 H atoms + 6 O atoms

Ionic Compounds

Ionic compounds are formed when metal atoms transfer electrons to nonmetal atoms resulting in the formation of ionic bonds.

Ionic compounds we use every day include:

- table salt NaCl
- baking soda NaHCO_3
- milk of magnesia $\text{Mg}(\text{OH})_2$

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David Meerman Scott

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