

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy

a. ionic bond

_____ bond energy

b. covalent bond

_____ enthalpy of vaporization

c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic

a. H—F

_____ nonpolar covalent

d. H—H

_____ ionic

b. Na—Cl

_____ polar covalent

e. H—C

_____ polar covalent

c. H—O

_____ polar covalent

f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Chapter 8 Covalent Bonding Worksheet Answers

Henry Sinclair Pickering



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Chemical Misconceptions Keith Taber, 2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources *Earth Science* Samuel N. Namowitz, Nancy E. Spaulding, 1989 **The Covalent Bond** Henry Sinclair Pickering, 1977

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In a world inundated with displays and the cacophony of immediate transmission, the profound power and psychological resonance of verbal artistry frequently diminish into obscurity, eclipsed by the continuous assault of sound and distractions. However, situated within the musical pages of **Chapter 8 Covalent Bonding Worksheet Answers**, a captivating function of fictional brilliance that impulses with organic thoughts, lies an unique journey waiting to be embarked upon. Written by a virtuoso wordsmith, this mesmerizing opus instructions visitors on a psychological odyssey, softly revealing the latent possible and profound impact embedded within the elaborate web of language. Within the heart-wrenching expanse of the evocative evaluation, we can embark upon an introspective exploration of the book is central styles, dissect their fascinating writing design, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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