

## Section 1 Review page 118

1. **Describe Democritus's atomic theory.**  
Everything in the universe is made up of atoms.
2. **Summarize the main ideas of Dalton's theory.**
  - a. Elements are made up of tiny unique particles called atoms.
  - b. Atoms cannot be divided.
  - c. Atoms of the same element are identical.
  - d. Atoms of different elements can join to form molecules.
3. **Explain why Dalton's theory was more successful than Democritus's theory.**  
Dalton had experimental evidence to support his theory, while Democritus did not.
4. **Compare Thomson's atomic model with Rutherford's atomic model.**  
Thomson's model- electrons were embedded in a sphere of positive charge  
Rutherford's model- positive charge is concentrated in the nucleus and electrons surrounded the nucleus.
5. **Analyzing elements**
  - a. **How did the cathode-ray tube experiment lead to the conclusion that atoms contain electrons?** The air was removed from the tube and the beam of negative particles was observed. This suggests that atoms are made up of smaller units with negative charges.
  - b. **How did the gold-foil experiment lead to the conclusion that the atom has a nucleus?** the positive particles were deflected at large angles and therefore this suggests that a positive charge is concentrated in the center of the atom. (nucleus)
6. **Making Inferences: Does the term *indivisible* still describe the atom? Explain.**  
*No*, the atom is not indivisible. It is made up of smaller parts such as the electron.

## Section 2 review page 127

1. **List the charge, mass, and location of each of the three subatomic particles found in atoms.** The proton is positive, the neutron is neutral and electrons are negative. Protons and neutrons are found in the nucleus of an atom. The mass of a proton 1850 times greater than that of an electron.
2. **Explain how you can use an atom's mass number and atomic number to determine the number of protons, electrons, and neutrons to determine the number of protons, electrons, and neutrons in the atom.** The atomic number equals the number of protons. The number of electrons equals the number of protons. To find the number of neutrons subtract the atomic number from the mass number.
3. **Identify the subatomic particle used to define an element, and explain why this particle is used.** The proton defines the element's number because the each element always has the same number of protons, while the number of neutrons and electrons can vary.
4. **Explain why the masses of atoms of the same element may differ.** Because they can have different number of neutrons. For example Hydrogen (H) H, H<sub>1</sub>Deuterium, H<sub>2</sub> Tritium. These are called isotopes.

# Chapter Review Physical Science

**G Thomas**



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