

## Study Guide Chapter 9 – Covalent Bonding

1. Describe covalent bonds and be able to distinguish between single, double, triple, sigma and pi bonds.
2. What determines the strength of covalent bonds?
3. Be able to draw Lewis structures (or electron dot diagrams) for Mg, Li, C, Al, S, Br<sub>2</sub>, MgO, HF, CH<sub>4</sub>, Li<sub>2</sub>S, and SO<sub>3</sub>. Name the structures and label ionic or covalent.
4. Be able to write and name molecular compounds and acids.
5. Draw the difference between simple formulas, Lewis structures, and structural formulas?
6. Give examples of exceptions to the octet rule including electron deficient molecules, expanded octets, and coordinated covalent bonds.
7. Describe hybridization and give the energy level diagram for carbon.
8. Be able to identify the seven different VSEPR models (or geometry). Build and draw the geometry of H<sub>2</sub>S, CO, AlF<sub>3</sub>, PCl<sub>3</sub>, MgBr<sub>2</sub>, PCl<sub>5</sub> and SBr<sub>6</sub>. State how many bonded pairs, lone pairs, and bond angles for each.
9. Draw SO<sub>3</sub> and describe why it is a resonance structure.
10. Can you discuss and draw the formation of an ester?
11. Draw methyl alcohol and formic acid then draw the ester and name it.
12. What are esters used for?
13. Draw a diagram of nonpolar covalent, polar covalent and ionic bonds. Distinguish between them with one sentence.
14. Be able to use electronegativities to distinguish between nonpolar, polar covalent and ionic bonds.
15. Describe how you would identify a molecules polarity.
16. Describe some properties of covalent compounds and solids.

# Chapter 9 Covalent Bonding Study Guide Answers

**Ronald Starkey**



## **Chapter 9 Covalent Bonding Study Guide Answers:**

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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*Study Guide* Joan I. Senyk, Darrell D. Ebbing, 1984

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