

Report Sheet

Name Megan Stokes

Partner(s) _____

Table 1. Titration Data.

	Trial 1	Trial 2	Trial 3
mass of Na_2CO_3 (grams)	.0793 g	.0849 g	.0914 g
moles of Na_2CO_3	.000748 mol	.000801 mol	.000862 mol
V final, volume dilute HCl at endpoint (mL)	19.13 mL	10.80 mL	11.15 mL
V initial, volume dilute HCl at start (mL)	11.15 mL	2.45 mL	2.38 mL
V total, volume dilute HCl to reach the endpoint (mL)	7.98 mL	8.35 mL	8.77 mL
molarity of dilute HCl in buret (mol/L)	.0938 M	.0959 M	.0983 M
average molarity of dilute HCl (mol/L)	.09601 M		
percent difference of trials (%)	2.396 %	1.146 %	2.385 %
Average molarity of <u>concentrated</u> unknown solution (mol/L)	4.8005 M		

Sample Calculations:

- What is the concentration (M) of the dilute HCl in the buret? Show an example calculation from one of your trials.

Trial 1: $\text{Na}_2\text{CO}_3 = .0793 \text{ g}$ $C_{\text{d HCl}} = \frac{n_{\text{Na}_2\text{CO}_3} \times \frac{1000 \text{ mL}}{1 \text{ L}}}{V_{\text{total}}} = \frac{.000748 \text{ mol} \times \frac{1000 \text{ mL}}{1 \text{ L}}}{7.98 \text{ mL}} = \boxed{.0938 \text{ M}}$

$\times \frac{1 \text{ mol}}{105.88 \text{ g}} = .000748 \text{ moles}$ $V_{\text{total}} = 19.13 - 11.15 = 7.98$

- Calculate the average concentration (M) of the dilute HCl from all three of your best titrations.

$$\frac{.0938 + .0959 + .0983}{3} = .09601$$

- Calculate the percent difference for one of the trials.

$$\% \text{ diff} = \left| \frac{\text{avg} - x}{\text{avg}} \right| \times 100 \%$$

Trial 1: $\left| \frac{.09601 - .0938}{.09601} \right|$

$$= .02396 \times 100 \%$$

$$= \boxed{2.396 \%}$$

Chem Fax Lab 6 Answers

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University (later "of the University of Illinois") University of Illinois (Urbana-Champaign campus), 1901 Circular of
Information University of Chicago, 1918 Catalogue University of Virginia, 1910 **Host Bibliographic Record for**
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- Chapter 3: Chem Fax Lab 6 Answers in Everyday Life
- Chapter 4: Chem Fax Lab 6 Answers in Specific Contexts
- Chapter 5: Conclusion

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