

Chemical Bonding and Molecular Structure

Name _____

Instructor _____ Section _____ Partner _____

Exercise 1. Determining VSEPR Geometry (Use balloons to complete the Table)
(1/2 pt. for each box)

VSEPR Type	Lewis Structure	Number of Atoms Bonded to Atom A	Number of Lone Pairs Around Atom A	Geometry of Electron Clouds Around A	Molecular Geometry of Atoms Around A (Molecular Shape)	Bond Angle B-A-B
AB ₂	B-A-B			linear	linear	180°
AB ₃	$\begin{array}{c} \text{B}-\text{A}-\text{B} \\ \\ \text{B} \end{array}$			trigonal planar	trigonal planar	120°
AB ₂ E	$\begin{array}{c} \text{B}-\ddot{\text{A}}-\text{B} \end{array}$			trigonal planar	bent (linear)	120°
AB ₄	$\begin{array}{c} \text{B} \\ \\ \text{B}-\text{A}-\text{B} \\ \\ \text{B} \end{array}$			tetrahedral	tetrahedral	109.5°
AB ₃ E	$\begin{array}{c} \text{B}-\ddot{\text{A}}-\text{B} \\ \\ \text{B} \end{array}$		2	tetrahedral	trigonal pyramidal	109.5°
AB ₂ E ₂	$\begin{array}{c} \text{B}-\ddot{\text{A}}-\text{B} \\ \\ \text{B} \end{array}$		4	tetrahedral	bent	109.5

Chemical Bonding Lab Answers

Cathy Duffy



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