

## Chapter

## 6

## CHEMICAL BONDING

## MULTIPLE CHOICE QUESTIONS

- Which of the following molecules has a coordinate covalent bond?  
(a) HCl (b) NaCl (c)  $\text{NH}_4\text{Cl}$  (d)  $\text{AlCl}_3$  [LHR 05]
- The angle formed in  $\text{sp}$  hybridization is:  
(a)  $120^\circ$  (b)  $180^\circ$  (c)  $109.5^\circ$  (d)  $107.5^\circ$  [GRW 06]
- Ionic compounds are mostly obtained by the combination of group:  
(a) 3 and 5 (b) 2 and 5 (c) 4 and 8 (d) 1 and 7 [GRW 07]
- Dipole moment of  $\text{CO}_2$  is:  
(a) 1.84 D (b) Zero D (c) 0.95 D (d) 2.2 D [GRW 07]
- In  $\text{sp}^2$  hybridization, the orbitals are oriented at an angle of:  
(a)  $109.5^\circ$  (b)  $120^\circ$  (c)  $180^\circ$  (d)  $0^\circ$  [LHR 07]
- Which of the following species has unpaired electron in anti bonding molecular orbitals?  
(a)  $\text{H}_2$  (b)  $\text{He}_2$  (c)  $\text{O}_2^{+2}$  (d)  $\text{N}_2^{+2}$  [MTN 08, 10, SWP, LHR 08]
- A molecular orbital can contain maximum electrons equal to:  
(a) One (b) Two (c) Three (d) Four [LHR 08]
- Carbon dioxide and methane have dipole moment value:  
(a) Zero and 1.85 D (b) 1.70 D and 1.80 D (c) Both have zero (d) None of these [GRW 08]
- Nobel gases have maximum stability and least reactive because:  
(a) They are very safe (b) Their valence shells are complete  
(c) They are gases (d) They are present in zero group [GRW 07]
- Which species has unpaired electron in anti-bonding molecular orbits?  
(a)  $\text{O}_2^{+2}$  (b)  $\text{N}_2^{+2}$  (c)  $\text{B}_2$  (d)  $\text{F}_2$  [FSD 11, GRW 09, 12]
- Octet rule is not followed in the formation of:  
(a)  $\text{NF}_3$  (b)  $\text{CF}_4$  (c)  $\text{CCl}_4$  (d)  $\text{PCl}_5$  [LHR 09]
- The most stable elements are:  
(a) Halogens (b) Lithium family (c) Noble gases (d) None of these [GRW 09, LHR 14]
- The hybridization of carbon in  $\text{C}_2\text{H}_4$  is:  
(a)  $\text{sp}$  (b)  $\text{sp}^2$  (c)  $\text{sp}^3$  (d) not hybridized [LHR 09]
- Positive ions are called:  
(a) Cations (b) Anions (c) Molecules (d) Hydrated ions [LHR 11]
- Total number of bonds in  $\text{C}_2\text{H}_4$  molecule is:  
(a) Six (b) Four (c) Five (d) Eight [LHR 10, 11]
- The S.I unit of dipole moment is:  
(a) Joule (b) Debye (c) Coulomb meter (d)  $\text{Nm}^{-2}$  [GRW 11]
- Ionic, covalent and co-ordinate covalent bond are present in:  
(a)  $\text{SO}_2$  (b)  $\text{NH}_4\text{Cl}$  (c)  $\text{C}_2\text{H}_2$  (d)  $\text{H}_2\text{O}$  [FSD, GRW 11]
- The highest electronegative element in the periodic table is:  
(a) Oxygen (b) Nitrogen (c) Chlorine (d) Fluorine [LHR 10]
- Bond angle between H-S-H bonds is:  
(a)  $105.5^\circ$  (b)  $107.5^\circ$  (c)  $92^\circ$  (d)  $95^\circ$  [LHR 10]
- S.I Unit of dipole moment is:  
(a) pm (b) Debye (c) mC (d) All [LHR 10]
- The amount of energy released by absorbing electron in the valence shell is:  
(a) Ionization energy (b) Electron affinity (c) Electro-negativity (d) Atomization energy [MTN 08, 09, GRW 10]
- The number of electrons shared in  $\text{SF}_6$ :  
(a) 4 (b) 12 (c) 6 (d) 8 [GRW 10]

# Chemical Bonding Section 1 Quiz Answers

**Michael B. Leyden**



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