

Chemistry--Unit 10: Chemical Periodicity  
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$  or abbreviated form\*:  $[\text{Ne}] 3s^2 3p^2$  \*may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$  or  $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an  $s^2 p^3$  configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a  $4s^2 4p^5$  electron configuration **Br, bromine**
- the 4<sup>th</sup> period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$  **O, oxygen**
- $4s^2$  **Ca, Calcium**
- $4s^2 3d^{10}$  **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

# Chemical Periodicity Chemistry Section Review Answers

**Williams & Wilkins Review**



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**Chemistry** Martin Stuart Silberberg,1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives

*Complete Preparation for the MCAT* Williams & Wilkins Review,1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT

*Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

*Modern Chemistry* Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

**Cehmistry Textbook for College and University USA** Ibrahim Sikder,2023-06-04 Cehmistry Textbook USA

A Complete Preparation for the MCAT Aftab S. Hassan,James L. Flowers,1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress

Chemistry John S. Phillips,Cheryl Wistrom,2000

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Study Guide to accompany Basic Concepts of Chemistry, 7th Edition Leo J. Malone, 2003-02-20. Work more effectively and gauge your progress along the way. This Study Guide, designed to accompany Malone's Chemistry 7th Edition, includes chapter summaries, new terms, self-tests, answers to self-tests, and solutions to selected problems. This easy-to-read introduction presents chemistry as a living, relevant science. Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject. This new seventh edition builds on its core strengths of pedagogy, driving the connections between ideas, mathematics in context, not just an appendix, and an extensive problem-solving emphasis with an updated design and more molecular art. In addition, the

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## **Table of Contents Chemical Periodicity Chemistry Section Review Answers**

1. Understanding the eBook Chemical Periodicity Chemistry Section Review Answers
  - The Rise of Digital Reading Chemical Periodicity Chemistry Section Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Periodicity Chemistry Section Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Periodicity Chemistry Section Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Periodicity Chemistry Section Review Answers
  - Personalized Recommendations
  - Chemical Periodicity Chemistry Section Review Answers User Reviews and Ratings

- Chemical Periodicity Chemistry Section Review Answers and Bestseller Lists
- 5. Accessing Chemical Periodicity Chemistry Section Review Answers Free and Paid eBooks
  - Chemical Periodicity Chemistry Section Review Answers Public Domain eBooks
  - Chemical Periodicity Chemistry Section Review Answers eBook Subscription Services
  - Chemical Periodicity Chemistry Section Review Answers Budget-Friendly Options
- 6. Navigating Chemical Periodicity Chemistry Section Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Periodicity Chemistry Section Review Answers Compatibility with Devices
  - Chemical Periodicity Chemistry Section Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Periodicity Chemistry Section Review Answers
  - Highlighting and Note-Taking Chemical Periodicity Chemistry Section Review Answers
  - Interactive Elements Chemical Periodicity Chemistry Section Review Answers
- 8. Staying Engaged with Chemical Periodicity Chemistry Section Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Periodicity Chemistry Section Review Answers
- 9. Balancing eBooks and Physical Books Chemical Periodicity Chemistry Section Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Periodicity Chemistry Section Review Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Periodicity Chemistry Section Review Answers
  - Setting Reading Goals Chemical Periodicity Chemistry Section Review Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Periodicity Chemistry Section Review Answers
  - Fact-Checking eBook Content of Chemical Periodicity Chemistry Section Review Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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