

**10****CHEMICAL QUANTITIES****Practice Problems**

*On a piece of paper, solve the following problems.*

***YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!***

**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

1. What is the molar mass of sucrose ( $C_{12}H_{22}O_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $PCl_5$ )
  - b. uranium hexafluoride ( $UF_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $KMnO_4$
  - b.  $Ca_3(PO_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $O_2$ )?

**SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS**

1. What is the molar mass of each of the following compounds?
  - a.  $C_6H_{12}O_6$
  - b.  $NaHCO_3$
  - c.  $C_3H_{12}$
  - d.  $KNH_4SO_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $PbO$ )
  - b. 0.75 mol hydrogen sulfide ( $H_2S$ )
  - c. 0.00100 mol silicon tetrahydride ( $SiH_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $O_2$ )
  - e. 2.30 mol ethylene glycol ( $C_2H_6O_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $N_2O_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $Ca_3(PO_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $NaBr$ )
  - b. 13.5 g magnesium nitrate [ $Mg(NO_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $MgCl_2$ )
  - d. 0.00100 g monochloromethane ( $CH_3Cl$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $C_3H_8(OH)_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $NH_4NO_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $C_{20}H_{24}N_2O_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $CH_4$ ) gas?
9. How many moles is 135 L of ammonia ( $NH_3$ ) gas at STP?

# Chemistry Chapter 10 The Mole Study Guide Answers

**Martha Mackin**



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remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

**Study Guide with Answers to Selected Problems** H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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**Fundamentals of General, Organic, and Biological Chemistry, Third Edition** John McMurry, Susan McMurry, 1999

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