## AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and Ka values:

HCIO<sub>4</sub> HOAc

HCN

HF

 $1 \times 10^7$   $1.76 \times 10^{-5}$   $4.93 \times 10^{-10}$   $3.53 \times 10^{-4}$  which shows the conjugate bases listed by increasing strength?

[A] ClO<sub>4</sub>-, F-, OAc-, CN-

[B] CN-, F-, OAc-, ClO<sub>4</sub>-

[C] CN-, OAc-, F-, ClO<sub>4</sub>-

[D] CN-, ClO<sub>4</sub>-, F-, OAc-

[E] ClO<sub>4</sub>-, OAc-, CN-, F-

Given that the K<sub>a</sub> for HOCl is 3.5 × 10<sup>-8</sup>, calculate the K value for the reaction of HOCl with OH<sup>-</sup>.

[A] 3.5 × 10-22

[B] 3.5 × 106

[C]  $2.9 \times 10^{-7}$ 

[D]  $3.5 \times 10^{-8}$ 

[E] none of these

 Acetic acid (HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>) is a weak acid (K<sub>a</sub> = 1.8 × 10<sup>-5</sup>). Calculate the pH of a 17.6 M HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> solution.

[A] 4.3

[B] 6.4

[C] 1.7

[D] 0.97

[E] 7.4

What concentration of acetic acid (K<sub>a</sub> = 1.80 × 10<sup>-5</sup>) has the same pH as that of 5.00 × 10<sup>-3</sup> M HCI?

[A]  $5.00 \times 10^{-3}$  M

[B] 3.60 × 10<sup>-3</sup> M

[C] 1.39 M

[D] 5.00 M

[E] none of these

Consider the reaction HOCl + F- - HF + OCl-

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**L** Cohen

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